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**LA THÈSE A ÉTÉ  
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SOLUBILITY OF THE TOXIC GASES,  
HYDROGEN SULFIDE AND SULFUR DIOXIDE IN  
SOME POLAR AND NON-POLAR SOLVENTS

by

Hassan Pahlevanzadeh

Thesis submitted to the School of Graduate  
Studies of the University Of Ottawa in  
partial fulfillment of the requirements  
for the degree of

MASTER OF APPLIED SCIENCE

in the

Department of Chemical Engineering

University of Ottawa

1985



H. Pahlevanzadeh , Ottawa , Canada , 1985

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ABSTRACT

The solubilities of the reactive gases, hydrogen sulfide and sulfur dioxide were measured in the following solvents: n-hexane, acetonitrile, dimethylformamide, N-N-dimethylacetamide and ethylacetate. In addition to the previously mentioned solvents, sulfur dioxide solubilities were also measured in acetone, methanol, chlorobenzene, and ethylene glycol. The solubility measurements were made at a total pressure of one atmosphere and at temperatures of  $-5^{\circ}\text{C}$ ,  $25^{\circ}\text{C}$  and  $60^{\circ}\text{C}$  where possible. The solubilities of sulfur dioxide were much higher than those for hydrogen sulfide in the same solvents and at the same temperatures. The solubility of sulfur dioxide in n-hexane was the highest obtained among the solvents used. In contrast, sulfur dioxide was least soluble in ethylene glycol and chlorobenzene. Finally, the highest, and lowest, solubilities of hydrogen sulfide were observed in dimethylacetamide, and n-hexane, respectively.

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## NOMENCLATURE

---

- C correction factor
- F solvent flow rate,  $\text{cm}^3/\text{min}$ .
- G Gibbs-Helmholtz free energy, J
- H heat of vaporization, J
- K Henry's constant
- L Ostwald coefficient
- P pressure, atm
- R gas constant, J/mole
- S entropy, J
- T temperature,  $^{\circ}\text{K}$  and  $^{\circ}\text{C}$
- V volume,  $\text{cm}^3$
- X mole fraction in liquid phase
- Y mole fraction in gas phase
- Z slope
- $\alpha, b, c$  parameters in equation (44)
- f fugacity, atm

## NOMENCLATURE, CONTINUED:

---

m	mole
$\alpha$	hydrogen-bonding factor
$\gamma$	activity coefficient
$\rho$	density, g/cm <sup>3</sup>
$\theta$	elapsed time, s
$\psi$	volume fraction
$\phi$	fugacity coefficient

\*\*\*\*\*

## SUPERSCRIPTS

---

G	gas phase
i	ideal
L	liquid phase
s	saturation
v	vapor
$\wedge$	mixture
o	pure
-	partial molal quantity
-	

NOMENCLATURE, CONTINUED:

---

SUBSCRIPTS

---

g gas

i component

$P_i$  partial pressure of  $i$ th component

r room temperature

s experimental temperature

t total

1 solvent

2 solute

## INTRODUCTION

In the chemical engineering design of absorption or separation equipment for chemical processes it is frequently necessary to have data for the solubilities of gases in liquids at various temperatures and pressures. This makes the equilibrium gas solubility an important subject to study. Numerous studies have been conducted based on industrial necessities. A completely reliable theory for predicting gas solubility in liquids has not been presented so far, although three models for predicting gas solubility have been reported. These models all have some limitations. These models include the regular solution theory and the concept of the ideal gas solubility as presented by Hildebrand [1], the cavity model theory as suggested by Uhlig [2], and the statistical thermodynamics model as suggested by Guggenheim [3]. Because of limitations in these theories, experimental research is still considered useful and necessary. Many experimental data for gas solubilities have been reported, but only a few of them are for highly soluble gases. The solubilities of trans-2-butene in carbon disulfide [4] and of n-butane in dodecane, heptane, hexane, and octane [5] corresponded to a concentration of nearly 80 mol% at  $-5^{\circ}\text{C}$ . The solubilities of isobutylene in hexane and ethylacetate were also approximately 60 mol% [6] at  $-5^{\circ}\text{C}$ . Finally, the solubility of ammonia in ethylene glycol at  $-10^{\circ}\text{C}$  was

observed to be 65 mol% [7]. Solubilities of toxic gases have been studied the least. Investigations on their solubilities are useful for separating and purifying sulfur dioxide and hydrogen sulfide from mixtures of gases (for example separating hydrogen sulfide from a mixture of gases coming out of oil wells or for manufacturing sulfuric acid).

In this work the solubilities of sulfur dioxide and hydrogen sulfide were measured in n-hexane and the polar solvents chlorobenzene, methanol, ethylacetate, dimethylformamide(DMF), acetone, N-N-dimethylacetamide(DMA), acetonitrile, and ethylene glycol. The measurements were conducted at a total pressure of one atmosphere and at temperatures ranging from  $-5^{\circ}\text{C}$  to  $60^{\circ}\text{C}$ .

BASIC THERMODYNAMICS OF VAPOR OR GAS - LIQUID EQUILIBRIA

Let us consider a liquid mixture which, at a temperature,  $T$ , and pressure,  $P$ , is in equilibrium with a vapor mixture at the same temperature and pressure. For every component  $i$  which exists in both mixtures, the condition of thermodynamics equilibrium is given by [8]

$$f_i^v = f_i^L \quad (1)$$

The fugacities are obtained as follows:

$$d\bar{G}^v = RT d \ln \hat{f}_i^v \quad (2)$$

$$d\bar{G}^L = RT d \ln \hat{f}_i^L \quad (3)$$

In the expressions above  $\bar{G}^v$  and  $\bar{G}^L$  are the partial molal Gibbs-Helmholtz free energies,  $f$  is the fugacity, the superscript  $v$  refers to the vapor, and superscript  $L$  refers to the liquid. The fugacity of a component in a mixture depends on the temperature, pressure, and composition of that mixture. The relationship of  $f_i^v$  and  $f_i^L$  to temperature, pressure and mole fraction is given below [8]:

$$f_i^v = \phi_i P_t y_i \quad (4)$$

$$f_i^L = \gamma_i x_i f_i^o \quad (5)$$

The fugacity coefficient of component,  $i$  ( $\phi_i$ ) in a gas mixture is a function of temperature only. The activity coefficient ( $\gamma_i$ ) at a certain temperature and pressure is defined by:

$$\gamma_i \equiv \frac{a_i}{x_i} \quad (6)$$

In the above expressions  $a_i$  is the activity of component  $i$ ,  $x_i$  and  $Y_i$  are compositions of component  $i$  in the liquid and vapor, respectively,  $f_i^\circ$  is the fugacity of component  $i$  in pure liquid and  $P_t$  is the total pressure. The fugacity coefficient ( $\phi_i$ ) is defined as follows [8] :

$$RT \ln \phi_i = RT \ln \frac{f_i^v}{P_t Y_i} = - \int_{\infty}^v \left\{ \left[ \frac{\partial P}{\partial N_i} \right]_{P, V, N_j [L]} - RT \right\} dV \quad (7)$$

For a mixture of ideal gases the fugacity coefficient is normalized such that as  $P \rightarrow 0$ ,  $\phi_i = 1$  for all components at low pressure. Substitution of equations (4), (5), and (7) in equation (1) gives:

$$Y_i P_t = \gamma_i x_i P_{VP_i} C_i \quad (8)$$

$$C_i = \frac{\phi_i^s}{\phi_i} \exp \int_{P_{VP_i}}^P \frac{V_i^L}{RT} dP \quad (9)$$

The term,  $C_i$  is a correction factor whose value is near unity when the total pressure is sufficiently low. If in addition to setting  $C_i=1$  we assume that  $\gamma_i=1$ , then equation (8) reduces to the familiar relation known as "Raoult's Law" :

$$P_i = X_i P_{v_i} \quad (10)$$

Based on equation ( 10 ) an ideal gas solubility can be defined for a binary gas-liquid system. For a gas partial pressure of one atmosphere, equation (10) reduces to an expression for the ideal gas solubility :

$$X_2^i = \frac{1}{p_2^o} \quad (11)$$

It is noted that the ideal solubility is dependent on temperature and the nature of the solute, but it is independent of solvent. Also, for a constant partial pressure the solubility of a gas decreases with increasing temperature. One of the useful characteristics of the ideal gas solubility relationship is that, when the system temperature is reduced to the normal boiling point of the solute gas, the ideal gas solubility becomes unity. The gas vapor pressure decreases with decreasing temperature to the normal boiling point of the solute gas. For this to actually occur, the solvent must not freeze at the gas

normal boiling point. The solubility of a gas varies greatly from solvent to solvent and in some instances it increases with increasing temperature, especially for gases of low solubility [4].

THE SOLUBILITY EQUATION BASED ON THE REGULAR SOLUTION THEORY

In 1929 Hildebrand [1] proposed the regular solution theory. This theory is meant to apply to solutions which are characterized by the absence of specific interactions between molecules. According to this theory, the change in Gibbs energy for transfer into the solution from the pure liquid state is as follows. [9] :

$$\Delta \bar{G}_2 = RT L_1 X_2 + V_2 \psi_1^2 (\delta_1 - \delta_2)^2 \quad (12)$$

where, 
$$\delta = \left( \frac{\Delta E^v}{V} \right)^{\frac{1}{2}} \quad (13)$$

In the above expression  $\delta$  is the solubility parameter and since it changes gradually with temperature, it should be calculated at the temperature at which the solubility is measured. The units for solubility parameter are  $(\text{cal}/\text{cm}^3)^{\frac{1}{2}}$ . Also in the above expression  $\psi_1$ , is a volume fraction for component 1 and  $V_2$  is the molar volume of pure liquid component 2. Because gas solubilities are relatively small, it is considered that the condition that  $\psi_1 = 1$  is valid. By using suitable thermodynamic relationships it can be shown that:

$$\Delta E^v = \Delta H^v - RT \quad (14)$$

By substitution of equation (14) into equation (13),  $\delta$  can

be written as :

$$\delta = \left( \frac{\Delta H^v - RT}{V} \right)^{\frac{1}{2}} \quad (15)$$

In the above expression,  $\Delta H^v$  is the heat of vaporization at the boiling point, and  $V$  is the liquid molar volume. On the other hand, on dissolution the Gibbs-free energy of a solute component is expressed in terms of the fugacities :

$$\Delta \bar{G}_2 = RT L_m \frac{f_2}{f_2^0} \quad (16)$$

When the vapor pressure is low or the gases are ideal, the fugacities can be replaced by the corresponding pressure  $f=P$  :

$$\frac{f_2}{f_2^0} = \frac{P_2}{P_2^0} \quad (17)$$

By inserting equation (16) into equation (12), the following is obtained:

$$L_m x_2 = L_m x_2^i - \frac{V_2}{RT} (\delta_1 - \delta_2)^2 \quad (18)$$

This last expression is considered very important for the prediction of solubilities of gases in liquids based on the

regular solution theory [10],[11], and [12]. It is not possible to estimate the solubility by means of equation (18) when the solubility is affected by a specific interaction between solute and solvent. The prediction of gas solubility by means of this equation is at least partially successful for a solution where both solute and solvent are non-polar or only slightly polar substances. It is evident from equation (18) and (11) that solubility of a gas is proportional to its vapor pressure (Henry's law). The ratio of the fugacities, rather than the ratio of vapor pressures, should be used in order to obtain the solubility more precisely, because the saturated vapor pressure of the gas  $P_2^0$  is usually large, and the temperature at which the solubility is sought is, in many cases, higher than the critical temperature of the gas.

The effect of a changing temperature on gas solubility was assessed by Hildebrand et al. [11]. This effect was obtained from the entropy of solution of the gas. The relationship between the Gibbs-free energy and mole fraction solubility at equilibrium is as follows:

$$\Delta G_2 = RT \ln x_2 \quad (19)$$

$$\Delta G_2 = \bar{G}_2^L - \bar{G}_2^G \quad (20)$$

$$x_2 = \frac{\bar{f}_2^L}{\bar{f}_2^G} \quad (21)$$

From thermodynamic relationships it is known that:

$$\left( \frac{\partial G}{\partial T} \right)_{P,x} = -S \quad (22)$$

The following relationship can be written for a solute component in solution:

$$\left( \frac{\partial \bar{G}_2^L}{\partial T} \right)_{P,x} - \left( \frac{\partial G_2^G}{\partial T} \right)_{P,x} = -\bar{S}_2^L + S_2^G \quad (23)$$

Substituting equation (19) in equation (23) results in:

$$\bar{S}_2^L - S_2^G = \left[ \frac{\partial (\bar{G}_2^L - G_2^G)}{\partial \ln x_2} \right]_{P,T} \cdot \left( \frac{\partial \ln x_2}{\partial T} \right)_{(\bar{G}_2^L - G_2^G), P} \quad (24)$$

Because the fugacity of gas,  $f_2^G$ , is not a function of composition ( $x_2$ ) equation (24) can easily be reduced to the following expression:

$$\bar{S}_2^L - S_2^G = R \left[ \frac{\partial \ln \hat{f}_2^L}{\partial \ln x_2} \right]_{P,T} \cdot \left( \frac{\partial \ln x_2}{\partial \ln T} \right)_{\text{Sat.}, P} \quad (25)$$

In the above expression the arbitrary standard state of the gas is taken as one atmosphere. When the pressure is low or

the solubility is low, the activity coefficient can be considered independent of  $X_2$  and the term,  $\frac{\partial \ln \hat{f}_2^L}{\partial \ln x_2}$  can be considered approximately equal to unity then:

$$\bar{S}_2^L - \bar{S}_2^G = \Delta \bar{S} = R \left[ \frac{\partial \ln x_2}{\partial \ln T} \right]_{\text{sat.}, P} \quad (26)$$

It may be observed from equation (26) that plots of  $\text{Log } X_2$  versus  $\text{Log } T$ , or plots of  $\text{Log } X_2$  versus  $\frac{1}{T}$ , are linear with slopes  $\Delta \bar{S}_2$ , and  $\Delta \bar{H}_2$ , respectively. In addition it may be observed from the above expression that the solubility of a gas is expected to increase with decreasing temperature.

The effect of molecular interaction on gas solubility in liquids is a strong function of the intermolecular forces between solute and solvent molecules. If  $x$  is a solvent and  $y$  a solute, the  $y$  molecules will dissolve in  $x$  only if  $x$ - $y$  interactions are energetically more favourable than the average of  $x$ - $x$  interactions. The intermolecular forces between chemically similar species lead to a smaller endothermic enthalpy of solution than those between dissimilar species. This interaction is "chemical" or "physical" affinity which sometimes causes enhancement in the solubility of gases in liquids. The solubility may be affected by the following considerations :

- (a) The occurrence of a reversible chemical reaction between the solute and solvent molecules, such

as that which takes place between sulfurous acid and water, or ammonia and water. In this case the extent of solubility is higher than the ideal gas solubility.

(b) The occurrence of hydrogen-bonding [13],[14]

between solvent-solvent molecules or solute-solvent molecules. For gases which tend to combine with the solvent by means of hydrogen-bonding or other complex formation, the solubility is enhanced when compared with that for a non-reacting gas. Conversely, if the hydrogen-bonding takes place between solvent molecules, then the solute molecules are excluded from penetrating into the solvent. Thus the solubility in such a solvent, with strong hydrogen-bonding tendencies, is less than that in a solvent with weaker hydrogen-bonds [16]. In solubility systems in which solvent hydrogen-bonding is predominant, the gas solubility is usually less than the ideal gas solubility.

(c) Existence of only a weak physical attraction

between solute and solvent molecules. In this case the intermolecular force is attributed to internal pressure or solubility parameter which is a function of the heat of vaporization. The resulting solubilities are normal or near the ideal gas solubility.

There are two sorts of solvents. One type has hydrogen attached to an electronegative atom and is known as a "protic solvent". Another type of solvent has no hydrogen atom attached to the electronegative atom. This type is known as an "aprotic solvent". These two kinds of solvents can enhance the gas solubility by forming hydrogen-bonding between solvent and solute molecules. An increase in electronegativity of an electronegative atom with an attached hydrogen atom makes the hydrogen-bond stronger and more stable. For example, hydrogen bond strengths in hydrogen sulfide are weaker than those in  $H_2O$  or  $HF$ . The solvents used in this work are : n-hexane, chlorobenzene, acetonitrile, methanol, ethylacetate, acetone, N-N-dimethylacetamide, dimethylformamide, and ethylene glycol. Among these solvents, only methanol and ethylene glycol are protic solvents because in their molecules there is a hydrogen atom attached to an oxygen atom. In a methanol molecule there is one hydrogen atom attached to a oxygen atom with a tendency to form a hydrogen bond with another negative atom. Thus, the methanol molecules in the liquid phase have hydrogen bonds among themselves. In that case, the attractive forces between the solute and methanol become small. This is also true for ethylene glycol which has two such hydrogen groups. Because the ethylene glycol has two hydrogen groups and methanol only one the solubility of gases used in ethylene glycol is expected to be lower than that in methanol. This was confirmed experimentally by

this work and Sahgals work[7]. The gases used in this work are hydrogen sulfide and sulfur dioxide. These gases are so reactive and polar as to be highly soluble in polar solvents because of the existence of chemical reaction or hydrogen-bonding between solute and solvent molecules.

For predicting gas solubility in a hydrogen-bonding or associating solvents when some solubility data are available, Hayduk and Laudie [16] have utilized a hydrogen-bonding factor based on the ideal solubility. This factor is defined as the ratio of the actual solubility to the ideal gas solubility, and is applicable for polar and hydrogen-bonding systems:

$$\alpha = \frac{X_2}{X_2^i} \quad (27)$$

$$X_2^i = \frac{P_2}{P_2^0} \quad (28)$$

It is assumed that hydrogen-bonding in a polar solvent causes solute gas molecules to be excluded or "squeezed out" from the liquid and hence the solubility of non-reactive gases is reduced. The solubility of a gas may be predicted from the relationship between hydrogen bonding-factors in different solvents, provided that the solubility is available in a reference solvent [16].

## EXPERIMENTAL APPARATUS AND PROCEDURE

---

Originally Morrisson and Billet [26] suggested a spiral tube-type of solubility apparatus for measurement of gas solubilities. This apparatus was modified by Hayduk and Cheng [17] and others; the resulting apparatus as used in this work is shown in figures 1, and 2. Figure 1 shows the solvent degassing apparatus and figure 2 the apparatus for solubility determinations. The procedure for degassing the solvent will now be described.

First about 30-40 cm<sup>3</sup> of solvent was put into the vacuum flask and a vacuum was applied. If the boiling point of the solvent was high, a tape heater was wrapped around the flask to raise the solvent temperature to its boiling point. After enough boiling had taken place for complete degassing of the solvent, (about 30, minutes) the stopcock was opened and the partially degassed solvent was allowed to flow into the packed bead column and degassing was continued for 10 minutes longer. Then the vacuum was stopped and the degassed solvent was temporarily stored in the packed column. The solvent was removed by means of a syringe using the septum rubber put at the end of the tube attached to the packed column. This solvent was subsequently slowly injected into the spiral tube of the solubility apparatus by means of a syringe driven by an infusion pump whose speed could be varied by changing its motor.

A 5-cm<sup>3</sup> gas-tight Hamilton syringe was used in

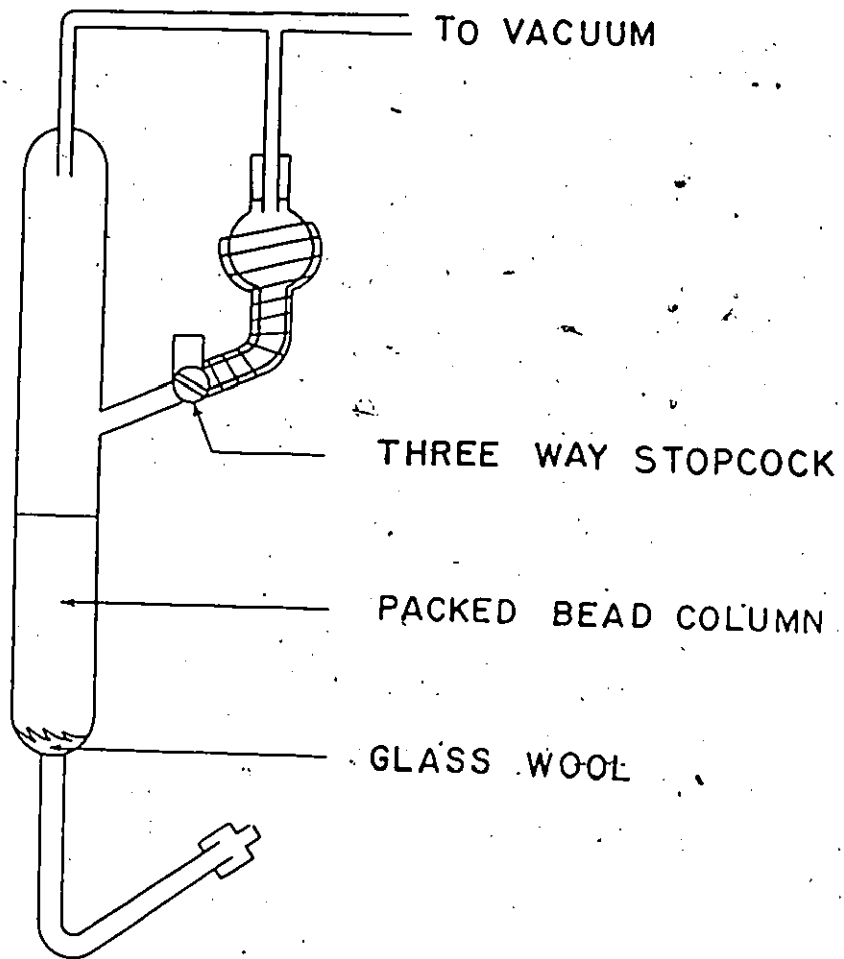


Figure 1  
Degassing apparatus

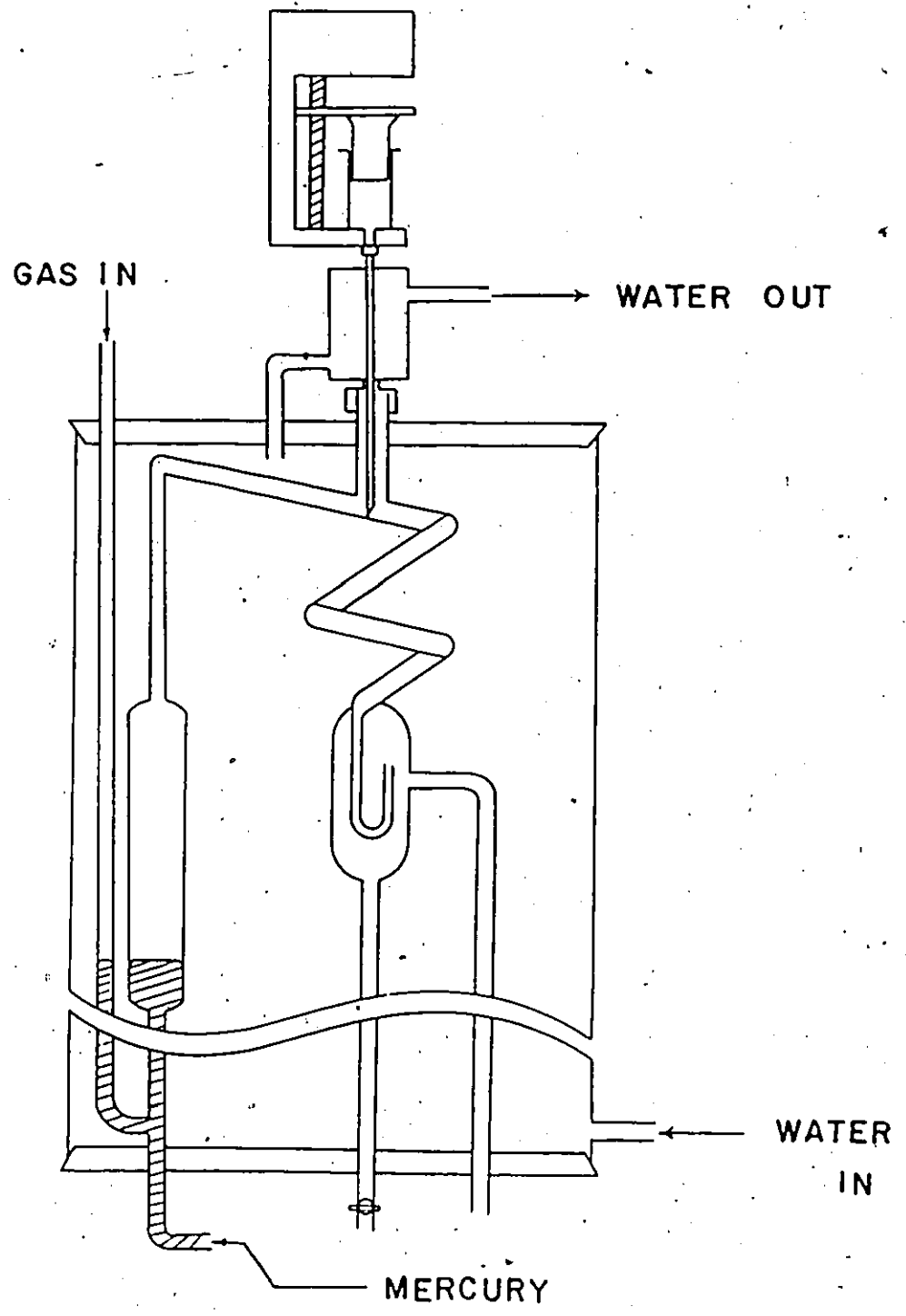


Figure. 2  
Solubility apparatus

conjunction with a syringe pump equipped with a motor chosen to give the necessary solvent flows. Three motor speeds, 0.5 rpm, 0.2 rpm, and 0.1 rpm were used to give the three different solvent flow rates of 0.04404 cm<sup>3</sup>/min., 0.01754 cm<sup>3</sup>/min., and 0.008687 cm<sup>3</sup>/min., respectively. These solvent flow rates were accurately measured by using distilled water at room temperature. For each combination of syringe and syringe-pump motor, at least three calibrations were made and the average of results was calculated. For calibration of the syringe pump, distilled water was injected into a closed vial with the needle penetrating the cap (to prevent evaporation) during a two to three hour period. The water so injected was weighed by means of a very accurate balance (10<sup>-5</sup> g). The solubility apparatus consisted of a gas burette, a spiral tube, a capillary U-tube monometer and a solution collecting burette. All of the components were of glass and were fixed inside a big tube which served as a jacket in which a water-glycol solution was circulated at constant temperature ( $\pm 0.05^{\circ}\text{C}$ ). To obtain a temperature of  $-5^{\circ}\text{C}$ , a Neslab Instruments refrigerating system was used. The temperature rise through the apparatus was less than  $0.5^{\circ}\text{C}$ ; to compensate for the temperature rise the inlet temperature was maintained at  $-5.25^{\circ}\text{C}$  and outlet temperature was  $-4.75^{\circ}\text{C}$ . Thus the average temperature inside the apparatus was maintained at  $-5^{\circ}\text{C}$ .

The top of the spiral tube was connected to the upper end of the gas burette. The tip of the long solvent supply

needle of the syringe was located at the top of the spiral tube. The lower end of the spiral tube was joined to the U-tube manometer. The solvent becomes saturated while flowing down the spiral tube to the U-tube. The pressure in the absorption spiral tube was observed by means of this U-tube, one side of which was open to the atmosphere. The inside pressure of the gas confined between the mercury surface and the surface of saturated solvent in the U-tube was controlled at atmospheric pressure. The size of the gas burette and spiral tube depended upon the gas solubility being measured. Three sizes of gas burette were used: 50 cm<sup>3</sup>, 100 cm<sup>3</sup>, and 250 cm<sup>3</sup>. The biggest size gas burette was used for high gas solubilities and the smaller sizes for lower solubilities. Several parameters were considered when selecting the design of spiral tube. First, when the solubility was high, the spiral tube was made longer because, in this case, the necessary residence time for saturating the solvent was considered to be long. The solvent residence time was the duration of time from the initial starting time of the syringe pump to the time when the first drop of liquid was observed in the U-tube manometer. The residence time for gases of low solubility was usually in the range of 10 -25 minutes. Because the flow rate of solvent was low, the solubility took place easily. Thus the required residence time for saturating the solvent was also low; therefore the necessary length of spiral tube for getting saturated solvent was also short.

This was confirmed experimentally by changing the solvent flow rate. Two different flow rates were used for the same solubility determinations in which the residence times were different. The solubility results, however, were found to be essentially identical. As the number of the coils in the spiral tube, the radius of the spiral coil, and the radius of the spiral tube itself were reduced in size, the residence time became shorter. The angle of descent of the spiral tube also had an effect on the residence time. By changing the slope of the spiral tube the residence time could be modified. In addition to the parameters mentioned, the solvent viscosity also influenced its residence time. At low temperatures the viscosity of solvent was higher, therefore the flow of liquid in the spiral tube was slower. If the slope of the spiral tube was too low, the liquid accumulated in the spiral tube and then suddenly moved. This phenomenon caused some fluctuations in the liquid levels in the U-tube. As a result, controlling the pressure of gas inside the apparatus became difficult. Hence, a large angle of descent for the spiral tube was used for high viscosity solvents.

For solubility measurements at temperatures higher than room temperature a heat exchanger was used to preheat the solvent to the experimental temperature. One heat exchanger was used to heat the long syringe needle. The syringe itself was also equipped with a heating ( or cooling ) jacket. The inlet water to the heat exchangers came from the

outlet of the column. In this way the heat exchanger temperature was adjusted to the system temperature. This method of preheating solvent in the syringe was not suitable for high temperatures because of the formation of some bubbles inside the syringe that affected the solvent flowrate. In addition, the calibration of the syringe for determining the liquid flow rate was seriously changed for temperatures which were higher or lower than room temperature. Hence, the syringe pump was calibrated at essentially room temperature, and the heat exchanger was used for the syringe needle only. The inner diameter of the long needle used was about 0.6 mm and the maximum solvent flow rate was  $0.01754 \text{ cm}^3/\text{min}$ . ; thus the maximum liquid velocity passing through the needle was about  $6.2 \text{ cm}/\text{min}$ . Because the length of the heat exchanger was about 10 cm , the residence time of liquid in the needle of about 1.6 minutes was sufficiently long for the solvent to essentially reach the experimental temperature.

The first steps in the procedure for making a solubility determination involved degassing the solvent and purging the solubility apparatus with solute gas. Then the syringe which had been cleaned well and completely dried was filled with about  $2-4 \text{ cm}^3$  of degassed solvent and connected to the infusion pump. The needle of the syringe was passed through the heat exchanger and then put into the spiral tube. The tip of the needle was extended into the apparatus until it touched the wall of the spiral tube for the purpose of

ensuring a continuous flow of solvent in the spiral tube. It was found that if the solvent flow to the spiral tube was dropwise, the inside pressure of gas inside the apparatus could not be controlled well. After closing the stopcock in the solution burette, the infusion pump was switched on. After the first droplet of solution appeared in the U-tube manometer, an additional waiting time was required for a steady-state condition to be achieved. Then the flow of solute gas was shut off and the mercury level was raised by means of the mercury-lift device driven by a variable-speed motor. The rate of the mercury-lift device could be adjusted as required so that the internal pressure was essentially atmospheric pressure as observed with the help of the U-tube manometer. The volume of mercury displaced was equal to the amount of solute gas dissolved in the spiral tube. This volume of mercury was recorded every five minutes; or instead, the elapsed time was recorded for every 10 cm<sup>3</sup> of mercury that had replaced dissolving gas. From the reduction in gas volume and solvent flow rate, the solubility was calculated.

## PROPERTIES OF TEST FLUIDS

Hydrogen sulfide and sulfur dioxide were purchased from the Matheson Company of Canada, and Air Product Chemical Company, respectively. The minimum purities of these gases were specified as 99.5 % , and 99.9 % respectively. The molar volumes for sulfur dioxide were obtained from the literature [ 18]. The molar volumes for hydrogen sulfide at a temperature of 25°C was obtained from reference [6], for a temperature of 60°C by extrapolating from the literature [19], and for a temperature of -5°C from another literature source [18]. These gas molar volumes are shown in table 1 .

The n-hexane, ethylacetate, dimethylacetamide, and acetone were purchased from Fisher Scientific Company. These solvents had specified minimum purities of 99+%, 99+%, 99+%, and 99.5% respectively. The chlorobenzene and ethylene glycol, both of which had a specified minimum purity of 99% , were supplied by the Aldrich Chemical Company. Dimethylformamide and methanol were obtained from J.T. Baker with specified purities of 99% , and 100% , respectively. Acetonitrile was obtained from Matheson with a specified purity of 99% . Vapor pressures of solvents were calculated by means of Antoine equation [8]. The densities of the solvents for a temperature of 25°C were obtained directly from the literature [2],[7], and [6]. The densities for a temperature of -5°C, for the solvents,

chlorobenzene, methanol, n-hexane, and ethylene glycol were also obtained from the literature, [20],[21],[22], and [23], respectively. For a temperature of 60°C, for chlorobenzene, n-hexane, and ethylene glycol, densities were obtained from [20] , [21], and [23] respectively. For the remaining solvents, densities were obtained by means of interpolating or extrapolating available data in the literature [6]. Yet another source of densities for a temperature of 20°C was [24]. At room temperature (ranging from 21°C to 27°C ) the densities of the solvents were considered to be a linear function of temperature:

## CHEMICAL FORMULA AND IMPURITIES OF SOLVENTS

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Solvent	Chemical Formula
Acetone	$\text{CH}_3\text{COCH}_3$
Acetonitrile	$\text{CH}_3\text{CN}$
Chlorobenzene	$\text{C}_6\text{H}_5\text{Cl}$
Dimethylformamide	$\text{HCON}(\text{CH}_3)_2$
N-N-Dimethylacetamide	$\text{CH}_3\text{CON}(\text{CH}_3)_2$
N-Hexane	$\text{C}_6\text{H}_{14}$
Ethylene glycol	$\text{HOCH}_2\text{CH}_2\text{OH}$
Ethylacetate	$\text{CH}_3\text{COOC}_2\text{H}_5$
Methanol	$\text{CH}_3\text{OH}$

TABLE 1

-----

PROPERTIES OF TEST FLUIDS

-----

gas molar volumes ,  $\text{m}^3/\text{kmol}$ .

-----

temp. K	$\text{SO}_2$	$\text{H}_2\text{S}$
268.15	21.423	21.723
298.15	24.047	24.264
333.15	27.021	27.181

\*\*\*\*\*

PROPERTIES OF SOLVENTS

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solvent	temp. K	density $\text{kg}/\text{m}^3$	vapor pressure mmHg
n-Hexane	268.15	682.7	34.46
	298.15	654.4	151.28
Ethylacetate	268.15	933.0	18.66
	298.15	894.2	96.81
	333.15	850.6	420.35
Chlorobenzene	268.15	1133	1.759
	298.15	1100.9	11.968
	333.15	1074.6	66.38

TABLE 1 (CONTINUED)

Solvent	Temp. K	Density Kg/m <sup>3</sup>	Vapor Pressure mmHg
Dimethyl acetamide	268.15	971.6	0.0
	298.15	936.8	0.0
	333.15	905.9	15.944
Dimethyl formamide	268.15	981.8	0.0
	298.15	944.3	0.0
	333.15	912.1	25.89
Acetonitrile	268.15	810.2	17.08
	298.15	777.3	86.33
	333.15	739.1	370.95
Methanol	268.15	814.7	21.76
	298.15	787.3	126.88
Ethylene glycol	268.15	1131.1	0.0078
	298.15	1110.1	0.127
	333.15	1085.6	0.647
Acetone	268.15	827.1	54.0
	298.15	785.3	30.65

### TREATMENT OF DATA

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The solvent flow rate was measured at room temperature but the experiments were performed not only at room temperature but also at other temperatures. The difference between room temperature ( syringe temperature ) and the system temperature caused a slight volume expansion or contraction of solvent in the apparatus; thus the solvent flow rate ,  $F_r$  , had to be corrected to obtain the actual liquid flow rate inside the solubility apparatus . This correction , to obtain the flow rate in the equipment ,  $F$  , is as follows:

$$F \cdot \rho_s = F_r \cdot \rho_r \quad (29)$$

$$F = F_r \cdot \frac{\rho_r}{\rho_s} \quad (30)$$

In the above equation  $\rho_r$  is the density of the solvent at room temperature and  $\rho_s$  is the density of the solvent at the experimental temperature. The linear relation between the absorption rate of dry gas, and the constant solvent rate to the apparatus was confirmed by results of all the experiments. The best slope for this linear relationship (solute gas volume- time) was calculated by means of a least-squares method. The Ostwald coefficient,  $L$  , is defined as the ratio of the volume of

the solute gas absorbed in the solvent to the volume of solvent. The Ostwald coefficient is virtually independent of pressure for small changes in pressure for which Henry's law holds. By applying Henry's law the mole fraction solubility of dissolved gas at a gas partial pressure,  $P_2$  can be calculated as shown below :

$$\frac{P_2}{X_{2,P_2}} = \frac{101.325}{X_2} = K \quad (31)$$

In the above expression  $X_2$  is the mole fraction of solute gas when the solute gas partial pressure is 101.325kPa and  $X_{2,P_2}$  is the mole fraction solute gas when the solute gas partial pressure is  $P_2$ . The equation for the gas partial pressure is then :

$$P_2 = \frac{X_{2,P_2}}{X_2} \cdot (101.325) \quad (32)$$

By considering that Raoult's law can be applied to the solvent for the purpose of estimating the partial pressure of solvent vapor  $P_i = X_i P_i^0$ ,  $P_2$  can be written as follows :

$$P_2 = 101.325 \left( \frac{1 - \frac{P_i}{P_i^0}}{X_2} \right) \quad (33)$$

$$P_2 = 101.325 \left( \frac{P_i^0 - P_i + P_2}{P_i^0 X_2} \right) \quad (34)$$

Hence, the expression for the partial pressure of gas is

reduced to:

$$P_2 = \frac{101.325(P_t - P_1^{\circ})}{101.325 - P_1^{\circ} x_2} \quad (35)$$

By definition, the Ostwald coefficient is first written in words and then expressed as an equation :

volume of gas (saturated) dissolved in solvent

---

volume of solvent

$$L = \frac{V_{w.g.}}{V_{\text{Solv.}}} \quad (36)$$

The volume of dry gas corresponding to wet gas dissolved in the solvent is expressed as :

$$V_g = \frac{P_2}{P_t} \cdot V_{w.g.} \quad (37)$$

Substituting equation (37) in equation (36) yields :

$$L = \frac{V_g}{V_{\text{Solv.}}} \times \frac{P_t}{P_2} \quad (38)$$

In the above equation the volume of solvent is the product

of the solvent flow rate and the time .

Thus :

$$L = \frac{V_g}{\theta \cdot F} \cdot \frac{P_t}{P_2} \quad (39)$$

In order to obtain the Ostwald coefficient from equation (39) and (35) the solubility needs to be known beforehand. From the definition of mole fraction of solute gas :

$$X_2 = \frac{m_2}{m_1 + m_2} \quad (40)$$

On the other hand , the equation for Ostwald coefficient is:

$$L = \frac{V_g}{V_{\text{Solv.}}} = \frac{m_2 V_2}{m_1 V_1} \quad (41)$$

In the above equation ,  $V_1$  , and  $V_2$  are the molar volumes of solvent and solute gas , respectively ; then it follows that :

$$\frac{m_1}{m_2} = \frac{V_2}{LV_1} \quad (42)$$

By inserting equation (42) into equation (40) we obtain:

$$X_2 = \frac{1}{1 + \frac{V_2}{LV_1}} \quad (43)$$

From equation (35) , (39) , and (43) the Ostwald coefficient

can be expressed as shown below:

$$L = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a} \quad (44)$$

The constants in the latter equation can be shown to be :

$$a = 101.325 (P_t - P_1^0) \quad (45)$$

$$b = 101.325 (P_t - P_1^0) \frac{V_2}{V_1} - \bar{P}_t Z (101.325 - P_1^0) \quad (46)$$

$$c = -101.325 P_t \cdot Z \cdot \frac{V_2}{V_1} \quad (47)$$

The slope of the absorption rate of (dry gas - solvent rate) relation,  $Z$ , is :

$$Z = \frac{V_g}{\theta \cdot F} \quad (48)$$

The value of  $Z$  was calculated by applying the least-squares method using the data obtained from the above experiments. The mole fraction solubility,  $X_2$ , was finally obtained by inserting equation (44) in equation (43). The basic computer program and the detailed results of the calculations are given in the Appendix.

## RESULTS AND DISCUSSION

Sulfur dioxide and hydrogen sulfide solubilities corresponding to a gas partial pressure of one atmosphere (101.325 kPa) and for temperatures of  $-5^{\circ}\text{C}$ ,  $25^{\circ}\text{C}$ , and  $60^{\circ}\text{C}$  in the solvents, n-hexane, acetone, methanol, dimethylformamide, N-N-dimethylacetamide, acetonitrile, ethylacetate, chlorobenzene, and ethylene glycol are listed in tables 3 and 5 as Ostwald coefficients. In tables 2 and 4 the results are listed as mole fractions and graphically indicated in figures 3 and 4. The solubilities were not obtained in methanol, n-hexane, acetone at  $60^{\circ}\text{C}$  because these solvents are above or near their normal boiling points at that temperature. Also, in tables 2 and 4 the average solubility is given for each system. For the purpose of confirming the accuracy of the experimental method, certain solubilities reported by Horiuti[25] were available for comparison. The solubilities of sulfur dioxide in chlorobenzene at  $25^{\circ}\text{C}$  and  $60^{\circ}\text{C}$ , as reported by Horiuti, are 0.182 and 0.0614 mole fraction, respectively; the values of Horiuti may be compared with those from this work which are listed in the same order and are, 0.1848 and 0.0617 mole fraction, respectively. The deviation between the two results are 1.5% and 0.33%, respectively. The solubility of hydrogen sulfide in chlorobenzene at  $25^{\circ}\text{C}$  was previously reported (Sahgal and Hayduk [7]) to be 0.0505 mole fraction. The solubility from this work is 0.05055, with a

TABLE 2

SULFUR DIOXIDE SOLUBILITY, MOLE FRACTION,  $X_2$ 

Solvent	Temp. K	1	2	3	Avg.
Ethylacetate	268.15	.8101	.8101		.8101
	298.15	.4351	.4397		.4374
	333.15	.220	.2145	.2196	.2198
Acetonitrile	268.15	.8096	.8120	.8077	.8098
	298.15	.3655	.3657	.3657	.3656
	333.15	.1518	.1545	.1524	.1529
Dimethyl- formamide	268.15	.8593	.8595	.8595	.8594
	298.15	.5875	.5916	.5927	.5906
	333.15	.3879	.3880	.3878	.3879
Dimethyl- acetamide	268.15	.8594	.8590	.8608	.8597
	298.15	.5969	.5901		.5935
	333.15	.4096	.4088	.4073	.4085
n-Hexane	268.15	.8989	.900	.9011	.900
	298.15	.6265	.6306	.6291	.6287
Acetone	268.15	.8350	.833	.835	.834
	298.15	.4494	.4488		.4491
Methanol	268.15	.699	.695	.698	.697
	298.15	.2584	.2568		.2576
Chloroben- zene	268.15	.7353	.739	.7365	.7369
	298.15	.1861	.1835		.1848
	333.15	.0616	.0618		.0617
Ethylene glycol	268.15	.535	.5354		.535
	298.15	.187	.1865	.1869	.1868
	333.15	.0671	.06763	.06629	.06701

TABLE 3

## SULFUR DIOXIDE SOLUBILITY , OSTWALD COEFFICIENT , L

Solvent	Temp. K	1	2	3	Avg.
Ethylacetate	268.15	968.2	967.9		968.05
	298.15	202.9	188	191.6	194.17
	333.15	71.3	73.4	73.6	72.77
Acetonitrile	268.15	1779.6	1826.2	1798	1800.3
	298.15	262.3	262.6		262.45
	333.15	87.1	88.9	87.5	87.83
Dimethyl- formamide	268.15	1761.6	1760.7	1758	1760.1
	298.15	450	452.1	442.4	448.2
	333.15	213.6	213.67	213.8	213.7
Dimethyl- acetamide	268.15	1432.5	1455.8	1460	1449
	298.15	383	372.4		377.7
	333.15	193	194	194.9	194
n-Hexane	268.15	1510	1546	1510	1533
	298.15	309.8	311.7	306.3	309
Acetone	268.15	1524	1547.9	1551	1540.97
	298.15	265.5	264.7		265.1
Methanol	268.15	1267	1241.7	1262.9	1257.2
	298.15	205.9	204.2		205

TABLE 3 (CONTINUED)

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SULFUR DIOXIDE SOLUBILITY , OSTWALD COEFFICIENT , L

-----

Solvent	Temp. K	1	2	3	Avg.
Chloroben- zene	268.15	602.9	607.5	610.6	607
	298.15	53.78	52.9		53.34
	333.15	16.93	16.99		16.96
Ethylene glycol	268.15	449.9	449.2		449.6
	298.15	98.9	98.6	98.8	98.8
	333.15	33.99	34.28	33.55	33.94

TABLE 4

 -----  
 HYDROGEN SULFIDE SOLUBILITY, MOLE FRACTION,  $X_2$   
 -----

Solvent	Temp. K	1	2	3	Avg.
n-Hexane	268.15	.0951	.0943	.09444	.0946
	298.15	.0370	.03656		.0368
Ethylacetate	268.15	.1853	.1874	.1840	.1855
	298.15	.08546	.08603		.08574
	333.15	.04035	.04041		.0404
Acetonitrile	268.15	.1030	.1026	.1022	.1026
	298.15	.0463	.04717		.04673
	333.15	.02098	.02077		.02088
Dimethyl- formamide	268.15	.2509	.2500	.2513	.2507
	298.15	.1162	.1150		.1156
	333.15	.05436	.05412	.0537	.0541
Dimethyl- acetamide	268.15	.2999	.3018	.3009	.3009
	298.15	.1457	.1456		.1456
	333.15	.06617	.06662		.0664
Chloroben- zene	298.15	.0506	.0505		.05055

TABLE 5

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 HYDROGEN SULFIDE SOLUBILITY , OSTWALD COEFF. , L  
 -----

Solvent	Temp. K	1	2	3	Avg.
Ethylacetate	268.15	53.33	53	51.88	52.74
	298.15	23.17	23		23.08
	333.15	11.03	11.05	13.23	11.77
Dimethyl- formamide	268.15	97.7	97.2	97.9	97.6
	298.15	40.7	41.23		40.97
	333.15	19.5	19.4	19.24	19.38
Dimethyl- acetamide	268.15	103.8	104.7	104.3	104.3
	298.15	44.5	44.5		44.5
	333.15	20	20.2		20.1
Acetonitrile	268.15	49.24	49.05	48.84	49.04
	298.15	22.3	22.7		22.5
	333.15	10.38	10.48		10.43
n-Hexane	268.15	17.9	17.9	18	17.9
	298.15	6.99	7.08		7.04
Chloroben- zene	298.15	12.65	12.64		12.645

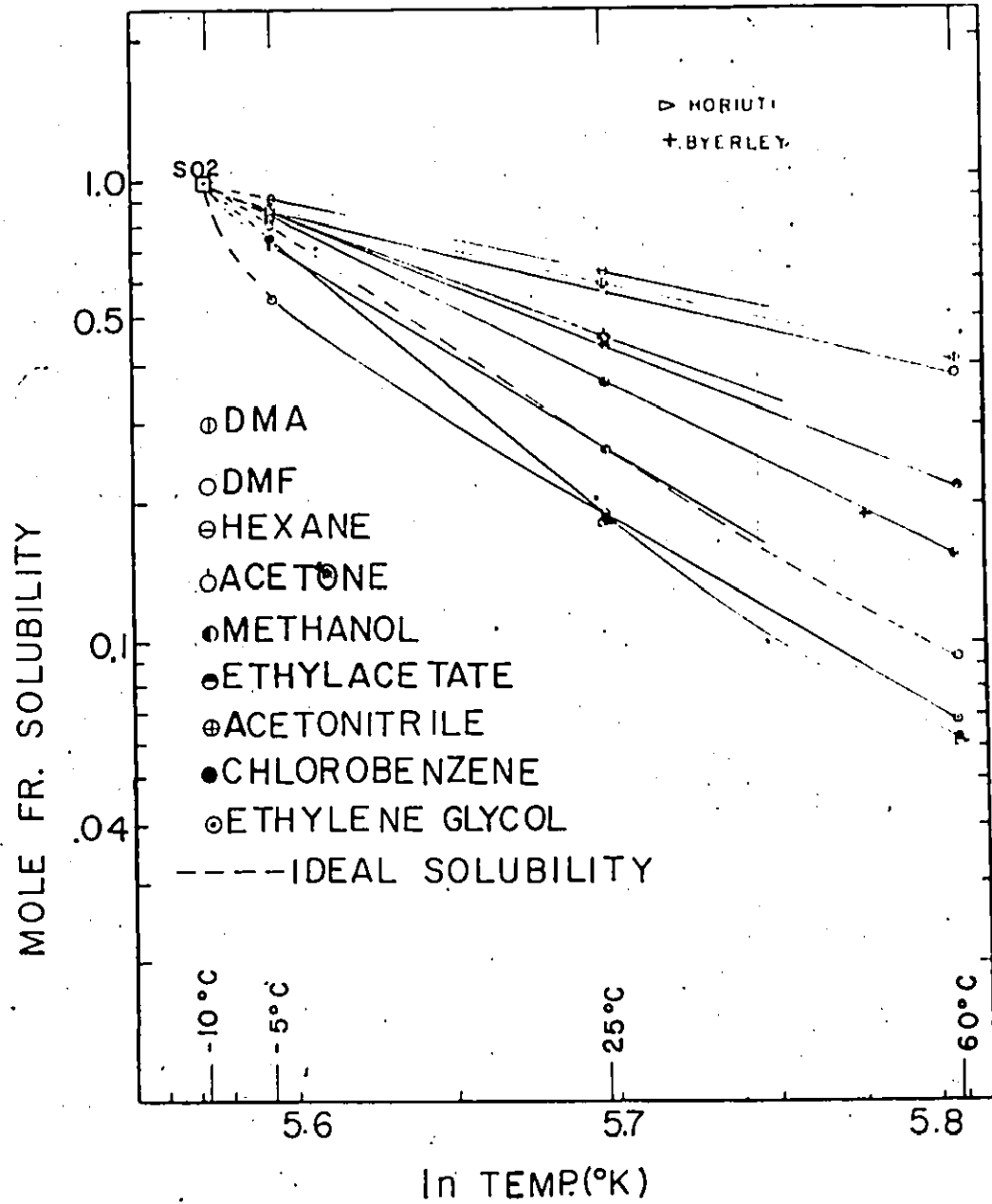


Figure 3

Solubilities of sulfur dioxide

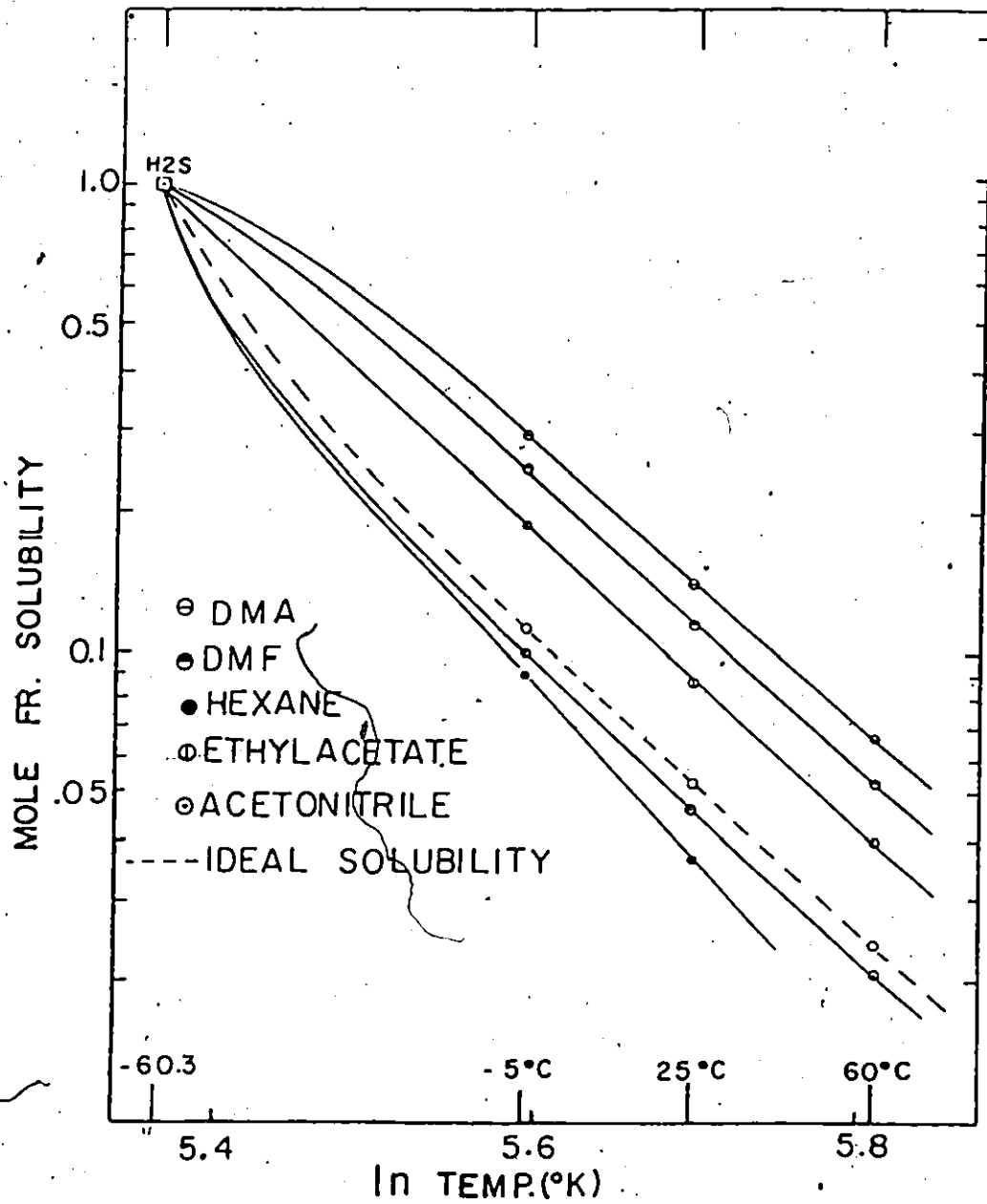


Figure 4

Solubilities of hydrogen sulfide

deviation of 0.1% . Solubility of sulfur dioxide in acetonitrile at 50 °C from this work as obtained by interpolating is 0.183 mole fraction while the value obtained by Byerley and Rempel [27] is 0.181 mole fraction. The deviation between the last two results is 1.1% . Also, the solubility of sulfur dioxide in acetone at 25 °C, as reported in the work of Horiuti is 0.451 mole fraction, and that from this work is 0.449 mole fraction. the deviation is 0.4% . All these comparisons show deviations from previously reported solubilities of less than 2% . It is considered that the apparatus used and method followed give results which are of good accuracy. Figure 4 shows that the solubility of hydrogen sulfide in n-hexane at 20 °C is about 0.041 mole fraction ; a comparable value was reported by Stephen and Stephen [29] as 0.0341 mole fraction. The deviation is large, with the previously reported value being 17% lower than that obtained in this work. Also, the solubility of sulfur dioxide in methanol at 25 °C from this work is 0.2576 mole fraction, while that from [29] is 0.201 mole fraction obtained by interpolation. The difference again is high, with the previously reported value being 22% lower than that obtained in this work. There is no apparent reason for these discrepancies . Common sources of error in the measurement of gas solubilities result in solubilities which are too low. In the case of the large deviations mentioned above, it is considered likely that the results previously reported are incorrect. In figures 3 and

4 are shown the gas solubilities which may be observed to increase with decreasing temperature ; in other words, the mole fractions solubility tend to become higher as the equilibrium temperature is reduced toward the boiling point of the gas. Large differences are observed for the solubilities of both gases in the various solvents. The results of the experiments show that the solubilities of sulfur dioxide are significantly higher in all solvents than the results for hydrogen sulfide. This is consistent with the observation that solubilities of gases increase in most solvents in the order of increasing normal boiling point. In figures 3 and 4 the ideal gas solubilities have been shown by a dashed line. In figure 3 it is observed that except for chlorobenzene and ethylene glycol, all the solubilities of sulfur dioxide are higher than the ideal solubility. This suggests that some strong association or chemical reaction occurs between sulfur dioxide and these solvents. It is difficult to explain the actual solubilities of sulfur dioxide in chlorobenzene and ethylene glycol. On the one hand chlorobenzene is only slightly polar, while on the other, ethylene glycol is highly self-associated.

The highest solubility of sulfur dioxide in the solvents used was obtained in n-hexane. Although n-hexane is a non-polar solvent, the solubility of sulfur dioxide in it is a maximum value among the solvents used. It has been postulated that for most gas solubilities above the ideal, a chemical association occurs between the gas and solvent.

Thus one may postulate that there is an association between sulfur dioxide and n-hexane which results in a usually high solubility. On the other hand, strong molecular self-association in the solvent often results in a reduced gas solubility. Thus the solubility of most gases is unusually low in ethylene glycol. There is a noticeable increase in sulfur dioxide solubility as the degree of molecular association in the solvent increases [28]. The solubility of the polar gas seems to increase as the polarity of solvent increases. These statements appear to be true for hydrogen sulfide as may be observed in figure 4. Also, as the polarity and association of solvent increases, so does the solubility of hydrogen sulfide. The solvents chlorobenzene, methanol, acetonitrile, ethylacetate, acetone, dimethylformamide, and N-N-dimethylacetamide are in the order of increasing gas solubility corresponding to figure 3 and are at the same time, in the order of increasing polarity and degree of association.

In figures 3 and 4 it can be observed that the solubilities have a general tendency to change with temperature in a way similar to that for the ideal gas solubility. It would also appear that as the temperature decreases toward the gas normal boiling point, the solubilities of most of these gases approach a mole fraction of unity. With an increase in temperature, sulfur dioxide is more soluble than hydrogen sulfide. This is probably due to the fact that sulfur dioxide is a more reactive and polar

gas. In table 6 the ratios of gas solubilities for sulfur dioxide to those of hydrogen sulfide in the same solvents and temperatures are shown. This comparison of solubilities may be used for choosing the best solvent to separate these two gases from each other. It should be noted, however, that there are several other considerations for selecting the best solvent. As table 6 shows, the largest ratio of solubility of sulfur dioxide to that of hydrogen sulfide occurs for the solvent ethylene glycol. Thus it can be concluded that ethylene glycol is the most effective solvent when compared with the other solvents used in this work for separating sulfur dioxide from hydrogen sulfide in the gas phase.

The hydrogen-bonding factor (H.B.F) is considered to be a measure of the effect of molecular association on gas solubility. The relation between the various H.B.F can be used to estimate solubilities in these and other associated solvents [16]. In figures 5 and 6 is shown the relation between H.B.F for a number of gases in ethylacetate, dimethylformamide, N-N-dimethylacetamide, acetone, ethylene glycol, and methanol at 25°C. The solvents dimethylformamide and acetone which are highly polar and associated were selected as reference solvents. The new data obtained in this research were converted into H.B.F and shown in these figures. From figures 5 and 6 it may be observed that for the solvents shown, except for the H.B.F of ethylene glycol and methanol for sulfur dioxide, the remainder are greater

TABLE 6

## RATIO SOLUBILITY OF SULFUR DIOXIDE TO HYDROGEN SULFIDE

Solvent	Temp. K	Ratio Solubility
n-Hexane	268.15	3.144
	298.15	4.65
Ethylacetate	268.15	4.367
	298.15	5.10
	333.15	5.44
Acetonitrile	268.15	7.89
	298.15	7.83
	333.15	6.07
Dimethylformamide	268.15	3.43
	298.15	5.10
	333.15	7.17
Dimethylacetamide	268.15	2.89
	298.15	4.074
	333.15	6.15
Chlorobenzene	298.15	3.66
	333.15	2.06
Acetone	298.15	6.43
Ethylene glycol	298.15	15.31
	333.15	9.33
Methanol	298.15	9.33

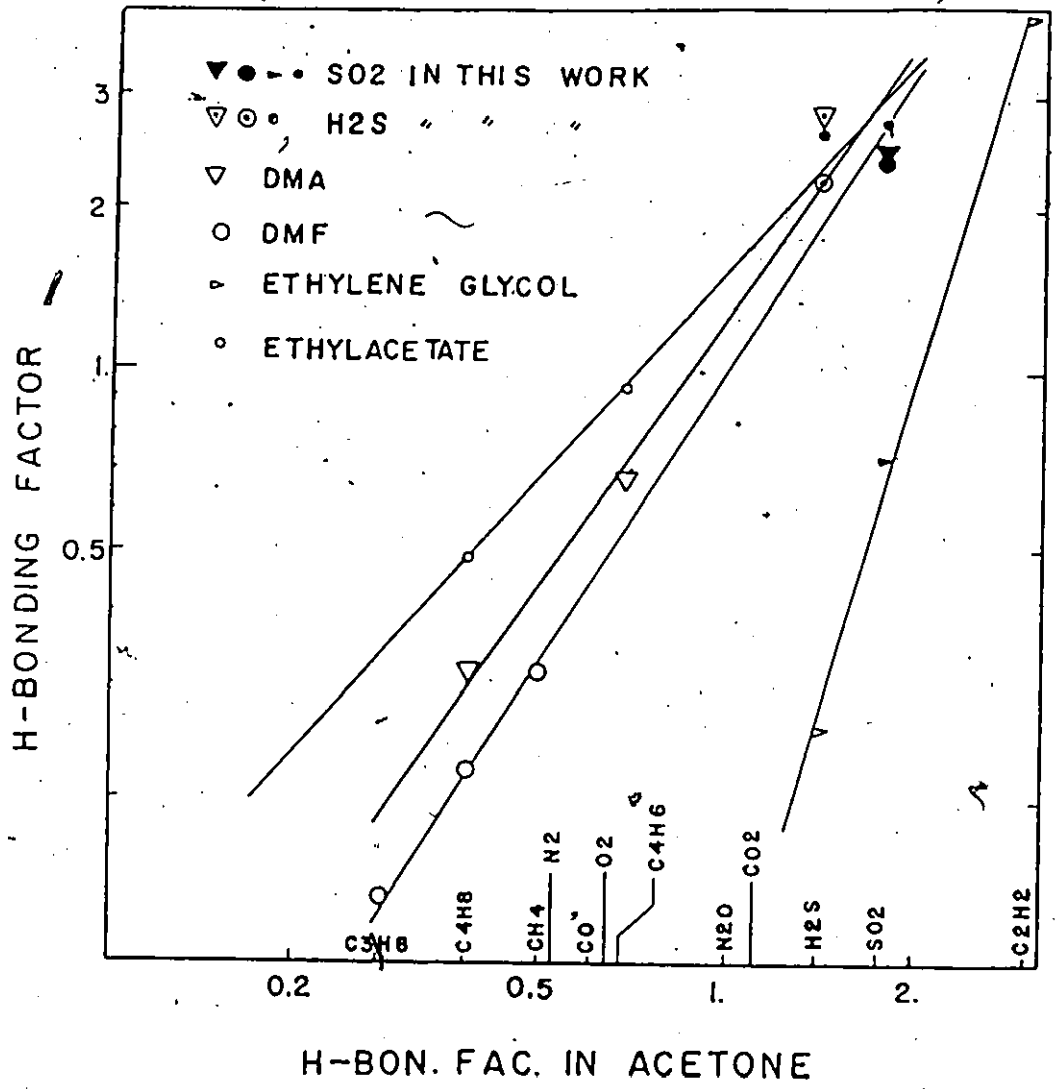


Figure 5  
Hydrogen-bonding factor correlation

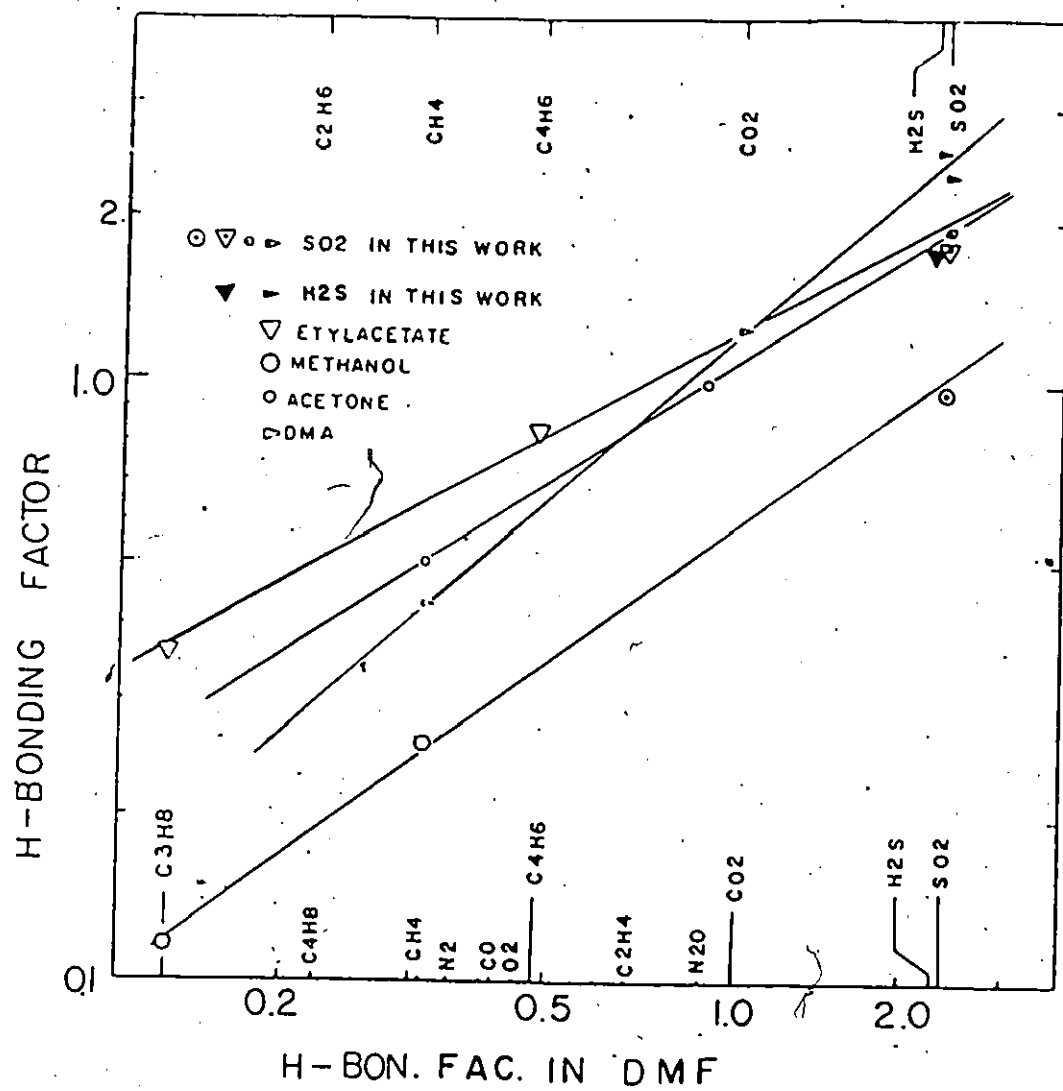


Figure 6

Hydrogen-bonding factor correlation

than one. Such high solubilities suggest that some chemical bonding occurs between these solvents and gases. These figures also indicate that it is possible to estimate the gas solubility from the relationship between H.B.F. in the solvents shown, provided that the solubility is known in at least one of them. Data for the H.B.F, shown in figures 5 and 6 are from sources given in table 7 ; the H.B.F obtained from this work are shown in tables 8 and 9.

The gases sulfur dioxide and hydrogen sulfide are toxic and perilous to human health. Hydrogen sulfide is both an irritant and an asphyxiant. At concentrations above 150 ppm the action of the gas on the nervous system becomes more pronounced . Exposure to 500 ppm results in a headache, and at 800-1000 ppm may be fatal in 30 minutes. Sulfur dioxide is dangerous to the eyes because it causes irritation and inflammation. It combines with water to form sulfurous acid. At a concentration of 6-12 ppm it causes immediate irritation of the nose and throat and with concentration of 400-500 ppm it is immediately dangerous to life. Maximum permissible concentration for exposures of 30-60 minute is 50-100 ppm [30].

For the above reasons all experiments involving hydrogen sulfide and sulfur dioxide were accomplished under a large fume hood for safety reasons. The gases discharged from the apparatus which had a high concentration of sulfur dioxide or hydrogen sulfide, were neutralized before coming out of the apparatus. An absorber solution was used at the

source of discharge of the gases for this purpose. The gases were diluted by means of dry air for improving the absorption. For absorbing sulfur dioxide, a solution of potassium permanganate ( 10 N ), and for hydrogen sulfide a solution of lead acetate(10% aqueous acetic acid solution) was used, both of which had a high capacity for absorbing these reactive gases.

TABLE 7

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SOLUBILITY DATA SOURCES FOR FIGURES 5, AND 6

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Solvent	Gases and Sources
Methanol	: $C_3H_8$ , $CH_4$ (27), $SO_2$ (This work)
Acetone	: $N_2O$ , $CH_4$ , $N_2$ , $C_2H_2$ , $O_2$ (23) $CO$ , $CO_2$ (Gjaldbaek and Anderson, 1954) $C_3H_8$ (19), $H_2S$ (6), $SO_2$ (This work)
Ethylacetate	: $C_4H_8$ , $C_4H_6$ (3), $C_2H_6$ (Mc Daniel, 1911) $SO_2$ , $H_2S$ (This work)
Dimethylformamide	: $C_2H_2$ , $C_2H_4$ , $C_3H_8$ , $CH_4$ , $CO_2$ (Howard et al., 1954) $C_4H_8$ (3), $SO_2$ , $H_2S$ (This work)
Ethylene glycol	: $C_2H_2$ (26), $H_2S$ (6), $SO_2$ (This work)

TABLE 8

-----

HYDROGEN-BONDING FACTORS FOR SULFUR DIOXIDE,  $\alpha$

-----

Solvent	Temp. K	$X_2$	$X_2^i$	$\alpha$
-----				
Ethylacetate	268.15	.8101	.8038	1.008
	298.15	.4374	.2584	1.69
	333.15	.2198	.09187	2.392
Acetonitrile	268.15	.8098	.8036	1.008
	298.15	.3656	.2584	1.42
	333.15	.1529	.09187	1.66
Dimethylformamide	268.15	.8594	.8036	1.07
	298.15	.5906	.2584	2.29
	333.15	.3879	.09187	4.22
Dimethylacetamide	268.15	.8598	.8036	1.07
	298.15	.5935	.2584	2.3
	333.15	.4085	.09187	4.45
n-Hexane	268.15	.9000	.8036	1.12
	298.15	.6287	.2584	2.43
Acetone	268.15	.834	.8036	1.038
	298.15	.4491	.2584	1.74

TABLE 8 (CONTINUED)

 -----  
 HYDROGEN-BONDING FACTORS FOR SULFUR DIOXIDE,  $\alpha$   
 -----

Solvent	Temp. K	$X_2$	$X_2^i$	$\alpha$
Methanol	268.15	.697	.8036	.867
	298.15	.2576	.2584	.997
Chlorobenzene	268.15	.7369	.8036	.917
	298.15	.1847	.2584	.72
	333.15	.0617	.09187	.67
Ethylene glycol	268.15	.535	.8036	.67
	298.15	.1868	.2584	.72
	333.15	.06701	.09187	.73

TABLE 9

-----

HYDROGEN-BONDING FACTORS FOR HYDROGEN SULFIDE,  $\alpha$

-----

Solvent	Temp. K	$X_2$	$X_2$	$\alpha$
-----				
n-Hexane	268.15	.2862	.1145	2.5
	298.15	.1351	.0512	2.64
Ethylacetate	268.15	.1855	.1145	1.62
	298.15	.08574	.0512	1.67
	333.15	.0404	.02445	1.65
Acetonitrile	268.15	.1026	.1145	.90
	298.15	.0467	.0512	.91
	333.15	.02517	.02445	1.03
Dimethylformamide	268.15	.2507	.1145	2.19
	298.15	.1156	.0512	2.26
	333.15	.0541	.02445	1.85
Dimethylacetamide	268.15	.3009	.1145	2.63
	298.15	.14565	.0512	2.84
	333.15	.0664	.02445	2.72
Chlorobenzene	298.15	.05055	.0512	.99

## CONCLUSIONS

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In this work it was found that the solubility of sulfur dioxide is much higher than that of hydrogen sulfide. In the same solvent and temperature it is higher by a factor which varies from 3 to 7.5. This result indicates that the tendency for sulfur dioxide to combine with the solvents used is greater than that for hydrogen sulfide in the same solvents and at the same temperatures. It was also found that the gas solubilities of the two gases studied had the maximum difference in ethylene glycol solvent. Thus this solvent was the most suitable absorbing solvent for separating sulfur dioxide from hydrogen sulfide. The plot of hydrogen-bonding factor in one solvent versus those in the other solvents was useful as a semi-empirical correlation for relating solubilities of sulfur dioxide and hydrogen sulfide in polar or associating solvents.

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## APPENDIX

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10     COMPUTER PROGRAM AND CALCULATION PRINTOUTS
20     *****
30     MAJOR PARAMETERS IN THE PROGRAM
40     TR = ROOM TEMPERATURE °C
50     TE = EXPERIMENTAL TEMPERATURE °C
60     V1 = SOLUTE GAS MOLAR VOLUME CM3
70     V2 = SOLVENT MOLAR VOLUME CM3
80     DR = DENSITY OF SOLVENT AT ROOM TEMP. G/CM3
90     DE = DENSITY OF SOLVENT AT EXPERIMENTAL TEMP. G/CM3
100    FA = ATMOSPHERIC PRESSURE KPA
110    PO2 = SOLVENT VAPOR PRESSURE KPA
120    FR = INFUSION RATE CM/MIN.
130    S = SLOPE
140    X = MOLE FRACTION SOLUBILITY
150    L = OSTWALD COEFFICIENT
200    CLS
210    ST=0
220    SV=0
230    STT=0
240    SVV=0
250    STV=0
260    INPUT "ENTER 3 NUMBERS : ",G,T,V
270    IF G=0 GOTO 430
280    N=G
290    IF G>9 GOTO 330
300    T(G)=T
310    V(G)=V
320    GOTO 360
330    M=G-9
340    F(M)=T
350    Q(M)=V
360    IF G<4 GOTO 210
370    ST=ST+T
380    STT=STT+T*T
390    SV=SV+V
400    SVV=SVV+V*V
410    STV=STV+V*T
420    GOTO 260
430    N=N-3
440    K=M
450    S=(N*STV-ST*SV)/(N*STT-ST*ST)
460    VO=(SV*STT-ST*STV)/(N*STT-ST*ST)
470    INPUT "ENTER 10 NUMBERS : ",TE,TR,FA,V1,DD,DE,D25,PO2,FR,M
480    INPUT "WHAT IS THE SYSTEM : ",SYS#
490    FA=FA*101.325/760
500    V2=M/DE

```

```

510 DR=D25+(25-TR)*DD
520 P02=P02*101.325/760
530 F=FR*DR/DE
540 A=101.325*(PA-P02)
550 B=101.325*(PA-P02)*V1/V2-PA*S*(101.325-P02)/F
560 C=-101.325*PA*S*V1/(F*V2)
570 L=(-B+(B^2-4*A*C)^.5)/(2*A)
580 X=1/(1+V1/(V2*L))
590 LPRINT " "
600 LPRINT SYS#
610 LPRINT " "
620 LPRINT TAB(5) "TR = " TAB(10) TR TAB(30) "TE = " TAB(35) TE
630 LPRINT TAB(5) "V1 = " TAB(10) V1 TAB(30) "V2 = " TAB(35) V2
640 LPRINT TAB(5) "DR = " TAB(10) DR TAB(30) "DE = " TAB(35) DE
650 LPRINT TAB(5) "P02 = " TAB(10) P02 TAB(30) "PA = " TAB(35) PA
660 LPRINT TAB(5) "FR = " TAB(10) FR TAB(30) "R = " TAB(35) R
670 LPRINT TAB(5) "S = " TAB(10) S TAB(30) "VO = " TAB(35) VO
680 LPRINT TAB(5) "L = " TAB(10) L TAB(30) "X = " TAB(35) X
690 LPRINT ""
700 LPRINT " TIME IN MIN. VOL. IN CC DEV. IN (%)"
710 LPRINT " 0 0 0"
720 FOR G=1 TO 9
730 VL=S*T(G)+VO
740 DEV.=(V(G)-VL)/VL*100
750 LPRINT TAB(4) T(G) TAB(25) V(G) TAB(45) DEV.
760 NEXT G
770 FOR H=1 TO K
780 VL=S*P(H)+VO
790 DEV.=(Q(H)-VL)/VL*100
800 LPRINT TAB(4) P(H) TAB(25) Q(H) TAB(45) DEV.
810 NEXT H
820 LPRINT " "
830 LPRINT "*****"
840 END

```

502 - ETHYLENE GLYCOL

TR = 24.8	TE = 25
V1 = 24047.26	V2 = 55.91298
DR = 1.11024	DE = 1.1101
POB = 1.693194E-02	PA = 100.805
FR = 2.923334E-04	R = 0
S = 2.892607E-02	VO = -.2622215
L = 98.94999	X = .1870391

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
146	4	.9849837
284	8	.5937213
424	12	-2.026942E-02
562	16	3.608375E-02
701	20	-.0747122
839	24	-2.812548E-02
977	28	5.18417E-03
1115	32	3.018098E-02
1253	36	4.962628E-02
1392	40	-7.171117E-03
1530	44	.0121218
1669	48	-3.204908E-02

\*\*\*\*\*

503 - ETHYLENE GLYCOL

TR = 24.1	TE = 25
V1 = 24047.26	V2 = 55.91298
DR = 1.11073	DE = 1.1101
POB = 1.693194E-02	PA = 99.80511
FR = 2.923334E-04	R = 0
S = 1.0288426	VO = 4.1114127
L = 98.62108	X = .1865334

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
144	4	-1.03717
284	8	-.9886841
421	12	-.2603253
559	16	-7.245056E-02
698	20	-.1035
836	24	-4.164205E-03
974	28	6.691094E-02
1113	32	3.001398E-02
1251	36	8.152078E-02
1390	40	5.053212E-02
1530	44	-4.035014E-02
1669	48	-5.598112E-02

\*\*\*\*\*

## SD2 - ETHYLENE GLYCOL

TR =	23.6	TE =	25
V1 =	24047.26	V2 =	55.91298
DR =	1.11108	DE =	1.1101
PO2 =	1.693194E-02	PA =	99.99178
FR =	2.923334E-04	R =	0
S =	2.892963E-02	VO =	-.1281985
L =	98.88728	X =	.1869427

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
107	3	1.102981
212	6	-3.131404E-02
316	9	-1.504855
420	12	-.1850354
523	15	-4.331152E-02
626	18	.1014997
730	21	4.123824E-02
834	24	3.503572E-03
938	27	-2.385621E-02
1041	30	4.153377E-02
1249	36	-1.362509E-02
1456	42	1.535076E-02
1664	48	-2.179511E-02

## SD2 - ETHYLENE GLYCOL

TR =	24.3	TE =	60
V1 =	27021.1	V2 =	57.17484
DR =	1.11059	DE =	1.0855
PO2 =	.2195819	PA =	97.55181
FR =	2.923334E-04	R =	0
S =	1.014694E-02	VO =	.1316301
L =	33.99881	X =	6.711127E-02

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
382	4	-.1936849
673	7	.5671538
870	9	.4523621
1168	12	.1396936
1465	15	2.065449E-02
1761	18	-2.214595E-03
2057	21	-1.853418E-02
2353	24	-3.077831E-02
2649	27	-4.029239E-02
2945	30	-4.790871E-02
3240	33	-2.341445E-02
3536	36	-3.116484E-02
3827	39	9.243037E-02

## SO2 - ETHYLENE GLYCOL

TR = 25	TE = 60
V1 = 27021.1	V2 = 57.17484
DR = 1.1101	DE = 1.0856
P02 = .2195819	PA = 99.65847
FR = 2.923334E-04	R = 0
S = 1.022727E-02	VO = .1272727
L = 34.28346	X = 6.763516E-02

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
287	3	-2.040816
577	6	-.4712443
870	9	-.2770041
1162	12	-9.460233E-02
1454	15	1.515936E-02
1747	18	3.157659E-02
2040	21	4.330648E-02
2334	24	9.474062E-03
2628	27	-1.683128E-02
2921	30	-3.789122E-03
3214	33	6.890049E-03
3508	36	-.0126187

\*\*\*\*\*

## SO2 - ETHYLENE GLYCOL

TR = 25.5	TE = 60
V1 = 27021.1	V2 = 57.17484
DR = 1.10975	DE = 1.0856
P02 = .2195819	PA = 100.2984
FR = 2.923334E-04	R = 0
S = .0100073	VO = 4.4816305
L = 33.5561	X = 6.629533E-02

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
289	3	-11.07811
580	6	-4.547835
885	9	-3.620626
1179	12	-2.282063
1474	15	-1.525664
1767	18	-.9057959
2061	21	-.505423
2355	24	-.2030216
2648	27	7.055807E-02
2940	30	.324069
3233	33	.5018031
3527	36	.6222421
3821	39	.7243778

\*\*\*\*\*

## S02 - ETHYLENE GLYCOL

TR = 27.2	TE = -5
V1 = 21423.7	V2 = 54.8749
DR = 1.10856	DE = 1.1311
P02 = 1.039914E-03	PA = 99.23184
FR = 1.464333E-04	R = 0
S = 6.456781E-02	VO = -.2982553
L = 449.9033	-X = .5353996

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
65	4	2.599569
128	8	.4214751
190	12	.2537474
252	16	.1700905
314	20	.1197729
377	24	.1321928
439	28	.167613
500	32	1.487059E-02
562	36	3.098201E-02
624	40	1.986028E-02
686	44	1.077768E-02
748	48	3.202059E-03

\*\*\*\*\*

## S02 - ETHYLENE GLYCOL

TR = 25.4	TE = -5
V1 = 21423.7	V2 = 54.8749
DR = 1.10982	DE = 1.1311
P02 = 1.039914E-03	PA = 100.3517
FR = 1.464333E-04	R = 0
S = 6.455023E-02	VO = -4.561843E-02
L = 449.2703	X = .5350493

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
61	4	-2.09507
127	8	-1.867701
186	12	.3283758
248	16	.2327961
311	20	-.1472971
373	24	-.1315609
435	28	-.1203246
496	32	8.978553E-02
558	36	7.392198E-02
620	40	6.123476E-02
682	44	5.085667E-02
745	48	-9.220673E-02

\*\*\*\*\*

## S02 - n-HEXANE

TR = 24.9	TE = -5
V1 = 21423.7	V2 = 127.6959
DR = .6545	DE = .6827
P02 = 4.595622	PA = 101.0983
FR = 1.464333E-04	R = 0
S = .2109776	VO = 2.897228
L = 1510.154	X = .9000128

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
95	20	-12.81644
188	40	-6.017298
281	60	-3.508964
373	80	-1.951028
465	100	-.9918865
556	120	-.1670395
648	140	.2788231
740	160	.6158472
835	180	.5229692
926	200	.8763572
1018	220	1.069288
1110	240	1.230628
1202	250	-2.531202

\*\*\*\*\*

## S02 - n-HEXANE

TR = 24	TE = -5
V1 = 21423.7	V2 = 126.2312
DR = .6554	DE = .6827
P02 = 4.595622	PA = 100.9784
FR = 1.464333E-04	R = 0
S = .2162916	VO = 1.049656
L = 1546.074	X = .9010848

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
90	20	-2.514634
182	40	-1.026185
273	60	-.1618432
365	80	4.883051E-03
458	100	-.1110825
550	120	-8.366202E-03
643	140	-8.931486E-02
734	160	.1203362
826	180	.1633149
920	200	-1.896308E-02
1013	220	-6.951785E-02
1105	240	-2.161195E-02
1151	250	-5.126927E-04

\*\*\*\*\*

## SO2 - n-HEXANE

TR = 24.9	TE = -5
V1 = 21423.7	V2 = 126.2312
DR = .6545	DE = .6827
P02= 4.595622	PA = 101.0983
FR = 1.464333E-04	R = 0
S = .2109776	VO = 2.897220
L = 1510.227	X = .898974

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
95	20	-12.81644
188	40	-6.017398
281	60	-3.508964
373	80	-1.951028
465	100	-.9918865
556	120	-.1670395
648	140	.2788331
740	160	.6158472
835	180	.5229692
926	200	.8763572
1018	220	1.069288
1110	240	1.230628
1202	250	-2.531202

\*\*\*\*\*

## SO2 - n-HEXANE

TR = 25	TE = 25
V1 = 24047.26	V2 = 131.6901
DR = .6544	DE = .6544
P02= 20.16901	PA = 99.27184
FR = 2.923334E-04	R = 0
S = 8.152603E-02	VO = .4157521
L = 306.3392	X = .6265321

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
118	10	-.3569595
242	20	-.7200283
363	30	-3.232547E-02
485	40	.1104333
731	60	-1.979021E-02
976	80	1.857149E-02
1222	100	-.0405338
1468	120	-7.989852E-02
1712	140	8.349435E-03
1957	160	2.363764E-02
2203	180	-9.773148E-03
2447	200	4.504097E-02
2693	220	1.575371E-02
2939	240	-8.645099E-03
3062	250	-1.937491E-02

\*\*\*\*\*

S02 - n-HEXANE

TR = 24.5	TE = 25
V1 = 24047.26	V2 = 131.6901
DR = .6549	DE = .6544
P02 = 20.16901	PA = 99.53848
FR = 2.923334E-04	R = 0
S = 8.314811E-02	VO = -1.789729
L = 311.6956	X = .6305791

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
127	10	14.02404
253	20	3.913689
378	30	1.213698
500	40	.5421114
622	50	.1434168
745	60	-.2586777
985	80	-.1337575
0	0	-100
1466	120	-8.774998E-02
1705	140	1.587166E-02
1946	160	-1.029862E-02
2187	180	-3.064382E-02
2427	200	-5.363177E-03
2666	220	5.315657E-02
2908	240	-2.066252E-03
3028	250	6.909657E-03

\*\*\*\*\*

S02 - n-HEXANE

TR = 24	TE = 25
V1 = 24047.26	V2 = 131.6901
DR = .6554	DE = .6544
P02 = 20.16901	PA = 99.40515
FR = 2.923334E-04	R = 0
S = 8.265758E-02	VO = .321969
L = 309.8246	X = .6291756

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
122	10	-3.903384
242	20	-1.599518
361	30	-.5349758
482	40	-.4056526
601	50	1.647976E-03
842	70	.1149107
1085	90	-6.052287E-03
1327	110	-7.802181E-03
1568	130	5.459757E-02
1810	150	4.521697E-02
2052	170	3.805374E-02
2415	200	2.999251E-02
2658	220	-1.174095E-02
2901	240	-.0464859
3021	250	-1.221164E-02

SO2 - ACETONE

TR = 23.1	TE = 25
V1 = 24047.26	V2 = 73.959
DR = .7870481	DE = .7853
P02 = 4.086331	PA = 100.6317
FR = 2.923334E-04	R = 0
S = 7.599522E-02	VO = -1.010946
L = 265.461	X = .4494737

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
142	10	2.245563
280	20	-1.320896
416	30	-1.970605
545	40	-1.0059
676	50	-.718449
805	60	-.2745888
1062	80	.3814794
1323	100	.4714813
1589	120	.2125671
1852	140	.1916464
2117	160	.0807269
2381	180	3.685508E-02
2646	200	-.036196
2911	220	-9.588576E-02
3175	240	-.1139803

\*\*\*\*\*

SO2 - ACETONE

TR = 24	TE = 25
V1 = 24047.26	V2 = 73.959
DR = .78622	DE = .7853
P02 = 4.086331	PA = 99.99178
FR = 2.923334E-04	R = 0
S = 7.569791E-02	VO = -.6743506
L = 264.78	X = .4403381

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
139	10	1.546976
277	20	-1.448568
408	30	-.6964415
538	40	-.1276578
670	50	-8.642731E-02
801	60	6.724744E-02
934	70	-3.926506E-02
1065	80	7.014423E-02
1329	100	7.187475E-02
1593	120	7.302845E-02
1857	140	7.385255E-02
2126	160	-.1618717
2388	180	-5.123479E-02
2651	200	-4.119856E-04
2914	220	4.120877E-02
3047	230	9.912561E-03

## S02 - ACETONE

TR = 25.4	TE = -5
V1 = 21423.7	V2 = 70.22126
DR = .786932	DE = .8271
P02= 7.199408	PA = 100.8184
FR = 1.464333E-04	R = 0
S = .2095899	VD = -3.699273
L = 1524.134	X = .8332141

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
123	20	-9.421471
170	30	-6.047459
217	40	-4.264404
263	50	-2.767014
308	60	-1.404045
402	80	-.6900523
496	100	-.2566714
592	120	-.3139842
685	140	9.307798E-02
776	160	.6653318
866	180	1.234155
960	200	1.262208
1053	220	1.382995
1146	240	1.48387
1253	250	-3.443924

\*\*\*\*\*

## S02 - ACETONE

TR = 26.9	TE = -5
V1 = 21423.7	V2 = 70.22126
DR = .785552	DE = .8271
P02= 7.199408	PA = 100.7517
FR = 1.464333E-04	R = 0
S = .2125113	VD = -39.43102
L = 1547.921	X = .8353552

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
114	20	-231.5379
373	40	.4124993
468	60	-4.039404E-02
563	80	-.2653121
657	100	-.1885096
749	120	.2172156
844	140	5.112126E-02
938	160	5.969748E-02
1032	180	6.638586E-02
1126	200	7.172193E-02
1221	220	-2.055353E-02
1315	240	-8.862028E-03
1363	250	-8.864206E-02

\*\*\*\*\*

SO2 - ACETONE

TR = 26.9	TE = -5
V1 = 21423.7	V2 = 70.22126
DR = .783552	DE = .8271
PO2 = 7.199408	PA = 100.7517
FR = 1.464333E-04	R = 0
S = .2125113	VO = -39.43102
L = 1551.831	X = .8357018

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
114	20	-231.5379
373	40	.4124993
468	60	-4.039404E-02
563	80	-.2653121
657	100	-.1885096
749	120	.2172156
844	140	5.112126E-02
938	160	5.969748E-02
1032	180	6.630586E-02
1126	200	7.172193E-02
1221	220	-2.055333E-02
1315	240	-8.862020E-03
1363	250	-0.864206E-02

\*\*\*\*\*

SO2 - METHANOL

TR = 24.5	TE = -5
V1 = 21423.7	V2 = 39.32982
DR = .78771	DE = .8147
PO2 = 2.901095	PA = 101.4717
FR = 1.464333E-04	R = 0
S = .1778828	VO = 2.831655
L = 1267.469	X = .6994137

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
51	10	-15.99234
103	20	-5.453372
154	30	-.7464036
262	50	1.13894
319	60	.7112507
432	80	.4053555
546	100	4.436171E-02
661	120	-.3423059
774	140	-.365037
888	160	-.4922926
1001	180	-.4932873
1114	200	-.494083
1216	220	.3937577
1329	240	.3185661
1386	250	.2497433

\*\*\*\*\*

## S02 - METHANOL

TR = 26.1	TE = -5
V1 = 21423.7	V2 = 39.32982
DR = .706398	DE = .8147
PO2 = 2.901095	PA = 99.85845
FR = 1.464333E-04	R = 0
S = .1738767	VO = 2.907258
L = 1241.741	X = .6950846

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
52	10	-16.30992
105	20	-5.501309
159	30	-1.812083
215	40	-.721645
328	60	.1020646
443	80	8.175853E-02
559	100	-.1042385
672	120	.2067444
787	140	.1801488
903	160	5.131493E-02
1019	180	-.0486688
1134	200	-4.171538E-02
1249	220	-3.603244E-02
1364	240	-.0312898
1421	250	5.560612E-03

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## S02 - METHANOL

TR = 26	TE = -5
V1 = 21423.7	V2 = 39.32982
DR = .78648	DE = .8147
PO2 = 2.901095	PA = 100.5117
FR = 1.464333E-04	R = 0
S = .1769131	VO = -2.76522
L = 1262.914	X = .6986561

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
59	10	30.33305
116	20	12.63357
174	30	7.075345
230	40	5.471922
342	50	-13.40349
456	80	2.686359
571	100	1.77895
686	120	1.182874
801	140	.7613683
914	160	.67115
1030	180	.3035683
1146	200	1.143014E-02
1261	220	-.1462148
1377	240	-.3504698
1435	250	-.4400715

S02 - METHANOL

TR = 23.4	TE = 25
V1 = 24047.26	V2 = 40.69859
DR = .788612	DE = .7873
P02 = 16.91594	PA = 100.5784
FR = 2.923334E-04	R = 0
S = .05243	VO = .2660093
L = 205.9656	X = .2584819

TIME IN SEC.	VOL. IN CC.	DEV. IN (%)
0	0	0
186	10	-.1795591
372	20	1.163552
566	30	.1957585
758	40	-1.986109E-02
950	50	-.148788
1142	60	-.234541
1334	70	-.2957211
1517	80	.2477283
1898	100	.2223561
2281	120	.1177838
2472	130	9.782227E-02
2664	140	4.32285E-02
3049	160	-.0781021
3241	170	-.1125907

\*\*\*\*\*

SU2 - METHANOL

TR = 23	TE = 25
V1 = 24047.26	V2 = 40.69859
DR = .78894	DE = .7873
P02 = 16.91594	PA = 98.79187
FR = 2.923334E-04	R = 0
S = 5.180167E-02	VO = .4896744
L = 204.2173	X = .2568514

TIME IN SEC.	VOL. IN CC.	DEV. IN (%)
0	0	0
190	10	-3.213236
377	20	-9.441989E-02
571	30	-.2275617
763	40	-3.585483E-02
957	50	-.127576
1149	60	-1.631153E-02
1341	70	6.330954E-02
1726	90	.1119552
2115	110	-4.561683E-02
2499	130	4.461077E-02
2887	150	-2.737694E-02
3080	160	-2.424606E-02
3272	170	8.985342E-03

\*\*\*\*\*

## SD2 - ACETONITRILE

TR =	25.4	TE =	60
V1 =	27021.1	V2 =	55.54459
DR =	.77662	DE =	.7391
PO2 =	49.45326	PA =	100.5251
FR =	.01754	R =	0
S =	.9005098	VO =	-1.534744
L =	08.91821	X =	.1545306

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	3.5	17.93226
10	7.5	.3968412
15	11.98	5.926162E-02
20	16.48	2.759936E-02
25	20.97	-3.814152E-02
30	25.47	-4.140983E-02
35	29.98	-1.034366E-02
40	34.51	.0706068
45	38.99	4.627944E-03
50	43.48	-2.470871E-02
55	48	1.396533E-02
60	52.5	7.913399E-03
65	56.99	-1.472381E-02
70	61.5	-1.538261E-03

\*\*\*\*\*

## SD2 - ACETONITRILE.

TR =	24.8	TE =	60
V1 =	27021.1	V2 =	55.54459
DR =	.7764	DE =	.7391
PO2 =	49.45326	PA =	102.3916
FR =	.01754	R =	0
S =	.9017858	VO =	1.428571E-02
L =	07.48203	X =	.1524189

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	4.55	.5921862
10	9.05	.1977006
15	13.55	6.593495E-02
20	18.05	-1.056703E-05
25	22.55	1.958613E-02
30	27.05	2.659768
35	31.6	-7.351711E-02
40	36.1	3.957861E-02
45	40.6	1.319345E-02
50	45.1	-7.924808E-03
55	49.6	-.0251968

\*\*\*\*\*

## SO2 - ACETONITRILE

TR = 24	TE = 60
V1 = 27021.1	V2 = 55.54439
DR = .779	DE = .7391
PO2 = 49.45326	FA = 102.2983
FR = .008786	R = 0
S = .4500728	VO = -4.198347E-03
L = 87.11148	X = .1518714

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	2.25	.1707019
10	4.5	7.717998E-02
15	6.74	.102102
20	9	3.040448E-02
25	11.24	-6.776335E-02
30	13.5	1.492193E-02
35	15.75	1.047639E-02
40	18	7.142472E-03
45	20.25	4.549307E-03
50	22.5	2.475376E-03
55	24.75	7.783084E-04
60	27	-6.357768E-04
65	29.25	-1.832325E-03
70	31.5	-2.857914E-03

\*\*\*\*\*

## SO2 - ACETONITRILE

TR = 22.5	TE = 25
V1 = 24047.26	V2 = 52.81480
DR = .78155	DE = .7773
PO2 = 11.50972	FA = 100.725
FR = 1.460333E-04	R = 0
S = 3.572911E-02	VO = -.2871987
L = 262.3924	X = .3657761

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
287	10	.3305399
570	20	-.3904393
850	30	-.2743928
1129	40	-.1272488
1408	50	-3.875756E-02
1687	60	2.032375E-02
1966	70	6.256746E-02
2246	80	1.954885E-02
2526	90	.0394256
2807	100	-4.409596E-03
3087	110	-7.781377E-03
3368	120	-4.034957E-02

502 -- ACETONITRILE

TR = 23.1	TE = 25
V1 = 24047.26	V2 = 52.81488
DR = .78053	DE = .7773
P02= 11.50972	PA = 100.9917
FR = 1.464333E-04	R = 0
S = .0356583	VO = -.5118156
L = 262.3321	X = .3655461

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
175	6	4.741529
350	12	.2624468
520	18	-.1691601
688	24	-8.781186E-02
855	30	7.996273E-02
1024	36	-6.336234E-03
1192	42	.0169692
1360	48	.0344395
1529	54	-1.801081E-02
1698	60	-5.993111E-02
1865	66	1.376949E-02
2033	72	2.569223E-02
2202	78	-9.946567E-03
2370	84	1.961883E-03

\*\*\*\*\*

502 - ACETONITRILE

TR = 23	TE = 25
V1 = 24047.26	V2 = 52.81488
DR = .7807	DE = .7773
P02= 11.50972	PA = 99.68514
FR = 1.464333E-04	R = 0
S = 3.564253E-02	VO = -1.073123
L = 262.5946	X = .365778

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
293	10	10.94207
586	20	.9417952
871	30	9.502531E-02
1152	40	3.233047E-02
1433	50	-5.24112E-03
1714	60	-3.027953E-02
1995	70	-4.815098E-02
2274	80	2.751153E-02
2555	90	7.172145E-03
2835	100	2.655735E-02
3116	110	1.000244E-02
3397	120	-3.789122E-03
3678	130	-1.546767E-02

## S02 - ACETONITRILE

TR = 25.5	TE = -5
V1 = 21423.7	V2 = 50.87021
DR = .77645	DE = .8102
P02 = 2.277146	PA = 101.2183
FR = 1.464333E-04	R = 0
S = .2482186	VD = 3.345938
L = 1776.637	X = .8077666

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
36	10	-18.57875
73	20	-6.828939
110	30	-2.120656
150	40	-1.426167
190	50	-1.004726
229	60	-.3123352
268	70	.1881892
347	90	.5856273
428	110	.380086
509	130	.2396665
590	150	.1369263
669	170	.3517272
752	190	-3.308639E-03
833	210	-5.330467E-02
915	230	-.2021675
995	250	-.1291969

\*\*\*\*\*

## S02 - ACETONITRILE

TR = 26.3	TE = -5
V1 = 21423.7	V2 = 50.67021
DR = .77509	DE = .8102
P02 = 2.277146	PA = 100.7517
FR = 1.464333E-04	R = 0
S = .2546999	VD = .842256
L = 1826.242	X = .8120065

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
37	10	-2.592514
75	20	.2770257
112	30	-.2.149766
151	40	1.776154
191	50	1.030648
231	60	.539682
473	120	-1.0842
553	140	-1.19384
633	160	-1.275567
704	180	-.0838022
0	0	-100
775	200	.8905279
856	220	.5184208
897	230	.3017581

\*\*\*\*\*

## SQ2 - ACETONITRILE

TR = 26.2	TE = -5
V1 = 21423.7	V2 = 50.67021
DR = .77526	DE = .8102
P02 = 2.277146	PA = 100.8184
FR = 1.464333E-04	R = 0
S = .2509244	VD = 2.030794
L = 1798.845	X = .8096881

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
37	20	76.75657
75	20	4.077311
112	30	-.4457669
151	40	.1994479
190	50	.5906107
230	60	.4294961
270	70	.3147299
350	90	.1621158
430	110	6.523918E-02
512	130	-.3862633
590	150	-.5.076564E-02
671	170	-.2353636
751	190	-.2493878
834	210	-.6160586
915	230	-.7022585
975	250	1.345017

\*\*\*\*\*

## SQ2 - DIMETHYLACETAMIDE

TR = 25.4	TE = -5
V1 = 21423.7	V2 = 92.52549
DR = .938376	DE = .9416
P02 = 0	PA = 100.5917
FR = 1.464333E-04	R = 0
S = .208599	VD = .8552717
L = 1432.479	X = .860853

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
46	10	-4.313781
93	20	-1.258851
140	30	-.196724
188	40	-.179388
236	50	-.1689873
283	60	.185694
331	70	.1408514
428	90	-.1504961
523	110	4.315244E-02
620	130	-.1433793
714	150	.1368753
810	170	.1057114
907	190	-2.871843E-02
1003	210	-3.811791E-02
1098	230	4.480794E-02
1195	250	-5.240782E-02

## S02 - DIMETHYLACETAMIDE

TR = 25	TE = -5
V1 = 21423.7	V2 = 89.66859
DR = .9368	DE = .9716
PO2 = 0	PA = 100.4317
FR = 1.464333E-04	R = 0
S = .2055505	VO = 3.361076
L = 1455.858	X = .8590253

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
45	10	-20.70318
90	20	-8.51128
136	30	-4.202144
182	40	-1.891683
228	50	+ .4511265
275	60	.1879212
370	80	.7369549
469	100	.2363054
567	120	7.656233E-02
664	140	.1097001
760	160	.263554
859	180	3.950196E-02
958	200	-.13902
1055	220	-9.046007E-02
1152	240	-6.463004E-02
1200	250	-8.654036E-03

\*\*\*\*\*

## S02 - DIMETHYLACETAMIDE

TR = 24.6	TE = -5
V1 = 21423.7	V2 = 89.66859
DR = .937224	DE = .9716
PO2 = 0	PA = 100.3251
FR = 1.464333E-04	R = 0
S = .2062855	VO = 3.787377
L = 1460.403	X = .8594024

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
45	10	-23.49021
88	20	-8.844367
133	30	-3.918039
180	40	-2.245323
217	50	2.983806
274	60	-.5133399
370	80	-.140495
467	100	-.122538
564	120	-.1101931
660	140	4.588461E-02
758	160	-9.476256E-02
853	180	.1397026
951	200	1.757358E-02
1048	220	1.157028E-02
1145	240	6.568068E-03
1194	250	-3.687605E-02

## SO2 - DIMETHYLACETAMIDE

TR = 22.4	TE = 25
V1 = 24047.26	V2 = 92.99958
DR = .939556	DE = .9368
P02 = 0	PA = 99.68514
FR = 1.464333E-04	R = 0
S = 5.625189E-02	VD = -1.602701
L = 383.0194	X = .5969816

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
182	10	15.80584
365	20	5.65666
546	30	3.054436
728	40	-1.655276
909	50	.9483781
1090	60	.4825624
1271	70	.1524494
1451	80	-2.348348E-02
1632	90	-.2221481
1813	100	-.3805198
1992	110	-.4083777
2355	130	-.6651488
2720	150	-.9262863
3084	170	-1.0927
3267	180	-1.192394
3302	190	3.181787

\*\*\*\*\*

## SO2 - DIMETHYLACETAMIDE

TR = 22.3	TE = 25
V1 = 24047.26	V2 = 92.99958
DR = .939662	DE = .9368
P02 = 0	PA = 99.85845
FR = 1.464333E-04	R = 0
S = 5.469768E-02	VD = -5.407905E-02
L = 372.3953	X = .5901958

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
185	10	-.6457134
370	20	-.9119129
556	30	-1.178708
738	40	-.775945
920	50	-.5327151
1101	60	-.2793181
1465	80	-9.742772E-02
1829	100	1.202537E-02
2192	120	.1308116
2555	140	.2158317
2915	160	.3829305
3276	180	.4825881
3639	200	.5071751
4062	220	-.9579544

## SO2 - DIMETHYLACETAMIDE

TR =	25	TE =	60
V1 =	27021.1	V2 =	96.17177
DR =	.9368	DE =	.9059
P02 =	2.125692	PA =	100.1118
FR =	2.923334E-04	R =	0
S =	5.762607E-02	VD =	.1797727
L =	193.0936	X =	.4073183

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
174	10	-2.025217
343	20	.2731708
513	30	.8676393
690	40	.1458193
864	50	6.264603E-02
1038	60	7.286602E-03
1386	80	-6.183613E-02
1733	100	-4.572493E-02
2080	120	-3.479125E-02
2427	140	-2.730577E-02
2773	160	1.445779E-02
3120	180	1.493089E-02
3294	190	-2.402302E-05
3467	200	1.532217E-02

\*\*\*\*\*

## SO2 - DIMETHYLACETAMIDE

TR =	26.1	TE =	60
V1 =	27021.1	V2 =	96.17177
DR =	.935634	DE =	.9059
P02 =	2.125692	PA =	100.1251
FR =	2.923334E-04	R =	0
S =	.057915	VD =	.4792647
L =	194.2966	X =	.4089185

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
168	10	-2.047063
338	20	-.2719426
508	30	.3341614
681	40	.2019366
854	50	.1228008
1028	60	-2.646699E-02
1373	80	4.301256E-03
1719	100	-3.513627E-02
2065	120	-6.141703E-02
2410	140	-3.885122E-02
2756	160	5.809269E-02
3102	180	-7.306161E-02
3275	190	-7.935496E-02
3438	200	.204902

\*\*\*\*\*

## 502 - ETHYLACETATE

TR =	25	TE =	60
V1 =	27021.1	V2 =	103.5822
DR =	.8942	DE =	.8506
P02 =	56.04206	PA =	101.3517
FR =	.008786	R =	0
S =	.3459998	VO =	.41
L =	73.59651	X =	.2200439

TIME IN MIN.	VOL. IN CC.	DEV. IN (%)
0	0	0
5	2.2	2.803804
10	3.9	.7752613
15	5.6	6.811964E-05
20	7.33	7.155815E-05
25	9.060001	7.36835E-05
30	10.79	7.070806E-05
35	12.52	7.617213E-05
40	14.25	7.361702E-05
45	15.98	7.161515E-05
50	17.71	6.461943E-05
55	19.44	7.849178E-05
60	21.17	7.207747E-05
65	22.9	7.496136E-05
70	24.63	6.96961E-05

\*\*\*\*\*

## 502 - DIMETHYLACETANIDE

TR =	25.6	TE =	60
V1 =	27021.1	V2 =	96.17177
DR =	.936164	DE =	.9059
P02 =	2.125692	PA =	100.1118
FR =	2.923334E-04	R =	0
S =	5.815553E-02	VO =	-.344945
L =	194.9899	X =	.4096796

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
177	10	-.5168213
347	20	.8317512
520	30	.3481098
693	40	.1080344
865	50	8.089119E-02
1038	60	-.0341426
1382	80	-3.248113E-02
1726	100	-3.148423E-02
2070	120	-3.081961E-02
2414	140	-3.035577E-02
2757	160	6.351875E-03
3100	180	3.490394E-02
3273	190	9.958467E-04
3445	200	-4.272443E-04

## SD2 - ETHYLACETATE

TR = 24.2	TE = 60
V1 = 27021.1	V2 = 103.5822
DR = .89516	DE = .8506
PO2 = 56.04206	PA = 101.3783
FR = .01754	R = 0
S = .6676008	VO = .6579835
L = 71.27534	X = .2145932

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	3.4	-14.91464
10	6.83	-6.871984
15	10.28	-3.673122
20	13.74	-1.927185
25	17.18	-.968418
30	20.63	-.2707501
35	24.08	.2330594
40	27.53	.6139454
45	30.97	.8794177
50	34.42	1.122218
55	37.87	1.32164
60	40.31	-.9913509
65	43.76	-.66293
70	47.21	-.3798998

\*\*\*\*\*

## SD2 - ETHYLACETATE

TR = 25	TE = 60
V1 = 27021.1	V2 = 103.5822
DR = .8942	DE = .8506
PO2 = 56.04206	PA = 101.405
FR = .01754	R = 0
S = .6892733	VO = .304562
L = 73.41221	X = .2196139

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	3.24	3.125459
10	6.6	6.1795548
15	10.02	-.1448682
20	13.47	-.8.088003E-02
25	16.92	-4.294199E-02
30	20.38	3.123112E-02
35	23.83	4.197448E-02
40	27.27	1.331874E-02
45	30.72	2.365498E-02
50	34.17	3.190538E-02
55	37.6	-1.453635E-02
60	41.04	-.028825
65	44.49	-1.842273E-02
70	47.95	1.133798E-02

\*\*\*\*\*

## S02 - ETHYLACETATE

TR =	23	TE =	25
V1 =	24047.26	V2 =	98.53164
DR =	.8966	DE =	.8942
P02 =	12.90707	PA =	101.7116
FR =	1.464333E-04	R =	0
S =	2.551563E-02	VO =	.5932591
L =	188.0062	X =	.4351367

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
385	10	-4.001005
769	20	-1.062485
1155	30	-.2122497
1545	40	-3.724618E-02
1936	50	1.697065E-02
2328	60	1.059962E-02
2720	70	6.038485E-03
3112	80	2.622673E-03
3504	90	-2.543131E-05
3896	100	-2.151443E-03
4288	110	-3.89084E-03

\*\*\*\*\*

## S02 - ETHYLACETATE

TR =	22.9	TE =	25
V1 =	24047.26	V2 =	98.53164
DR =	.8967201	DE =	.8942
P02 =	12.90707	PA =	99.59181
FR =	2.923334E-04	R =	0
S =	5.178647E-02	VO =	.4386819
L =	191.5841	X =	.439776

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
190	10	-2.705853
381	20	-.8395262
573	30	-.3730286
766	40	-.2670766
956	50	.1070251
1149	60	-9.787275E-02
1343	70	1.727811E-02
1536	80	2.163401E-02
1729	90	.0250222
1923	100	-.0240497
2116	110	-1.712852E-02
2309	120	-1.136015E-02

\*\*\*\*\*

S02 - ETHYLACETATE

TR = 22.5	TE = -5
V1 = 21423.7	V2 = 94.43408
DR = .8972001	DE = .933
P02 = 2.487795	PA = 101.445
FR = 1.464333E-04	R = 0
S = .1356946	VD = .1857815
L = 968.2151	X = .8101682

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
70	10	3.258793
145	20	.6973204
220	30	-.1284981
294	40	-.1976054
368	50	-.242225
513	70	.2906613
661	90	.1335909
808	110	.1574825
957	130	-3.501253E-02
1107	150	-.2657713
1251	170	3.545784E-02
1399	190	-1.186832E-02
1546	210	1.444708E-02
1693	230	3.620297E-02

\*\*\*\*\*

S02 - ETHYLACETATE

TR = 22.1	TE = -5
V1 = 21423.7	V2 = 94.43408
DR = .8976801	DE = .933
P02 = 2.487795	PA = 101.565
FR = 1.464333E-04	R = 0
S = .135731	VD = -.4816661
L = 967.9289	X = .8101227

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
69	10	12.56482
151	20	-6.850316E-02
225	30	-.1922972
298	40	8.467208E-02
372	50	-2.050361E-02
445	60	.1358285
519	70	5.330337E-02
593	80	-8.496516E-03
668	90	-.2069294
814	110	-3.037794E-03
961	130	3.402694E-02
1109	150	-2.930863E-02
1256	170	2.100372E-03
1403	190	2.691089E-02

\*\*\*\*\*

S02 - CHLOROBENZENE

TR = 23.5	TE = 25
V1 = 24047.26	V2 = 102.2427
DR = 1.10237	DE = 1.1009
P02 = 1.595602	FA = 100.1518
FR = 2.923334E-04	R = 0
S = 1.553991E-02	VO = 2.149876E-02
L = 53.78862	X = .1861285

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
664	10	-3.288203
1304	20	-1.40762
1939	30	-.5086957
2578	40	-.208049
3215	50	3.538239E-02
3857	60	6.848253E-02
4499	70	9.213891E-02
5145	80	3.208236E-02
5790	90	2.678837E-03
6438	100	-6.740601E-02

\*\*\*\*\*

S02 - CHLOROBENZENE

TR = 22.6	TE = 25
V1 = 24047.26	V2 = 102.2427
DR = 1.103252	DE = 1.1009
P02 = 1.595602	FA = 99.24517
FR = 2.923334E-04	R = 0
S = 1.528414E-02	VO = 1.095536
L = 52.87101	X = .183536

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
694	10	-14.54982
1330	20	-6.64429
1952	30	-3.007317
2576	40	-1.155171
3196	50	.1128592
3832	60	.5625726
4496	70	.2678424
5150	80	.239525
5816	90	1.325995E-02
6480	100	-.1365473
7137	110	-.1619314

\*\*\*\*\*

SO2 - CHLORO BENZENE

TR = 24	TE = 60
V1 = 27021.1	V2 = 104.745
DR = 1.10188	DE = 1.0746
PO2 = 8.849938	PA = 101.045
FR = .01754	R = 0
S = .2793636	VQ = -3.226446E-02
L = 16.93237	X = 6.159405E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.42	4.063326
10	2.81	1.761012
15	4.2	1.005487
20	5.59	.6299221
25	6.97	.2614233
30	8.36	.1860149
35	9.74	-5.605323E-02
40	11.12	-.1999653
45	12.5	-.3118146
50	13.89	-.3294836
55	15.27	-.4091547
60	16.66	-.4157515
65	18.4	1.509559
70	19.43	-.4773259

\*\*\*\*\*

SO2 - CHLORO BENZENE

TR = 24.7	TE = 60
V1 = 27021.1	V2 = 104.745
DR = 1.101194	DE = 1.0746
PO2 = 8.849938	PA = 101.285
FR = .01754	R = 0
S = .2802913	VQ = 3.507438E-02
L = 16.99499	X = 6.180777E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.4	-2.542997
10	2.83	-.2814578
15	4.24	1.311474E-02
20	5.64	-1.596811E-02
25	7.04	-3.347577E-02
30	8.439999	-4.516613E-02
35	9.850001	4.804566E-02
40	11.24	-5.981491E-02
45	12.65	1.435643E-02
50	14.06	7.378686E-02
55	15.46	5.762377E-02
60	16.85	-.0151433
65	18.25	-2.196365E-02
70	19.65	-2.781141E-02

\*\*\*\*\*

S02 -CHLOROBENZENE

TR = 24.1	TE = -5
V1 = 21723.7	V2 = 99.34598
DR = 1.101782	DE = 1.133
PO2 = .234514	PA = 101.7116
FR = 1.464333E-04	R = 0
S = 8.646740E-02	VO = 1.157059
L = 607.5894	X = .7353517

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
225	20	-2.974044
453	40	-.8124091
685	60	-.76426444
915	80	-.3433131
1145	100	-.1620508
1375	120	-4.217911E-02
1604	140	.1060519
1834	160	.1832682
2064	180	.2078064
2294	200	.2434731
2535	220	-.1601546
2765	240	-.1600737
2880	250	-.0736152

\*\*\*\*\*

S02 - CHLOROBENZENE

TR = 24	TE = -5
V1 = 21423.7	V2 = 99.34598
DR = 1.10108	DE = 1.133
PO2 = .234514	PA = 101.8316
FR = 1.464333E-04	R = 0
S = 8.691521E-02	VO = .7694506
L = 610.6737	X = .7390271

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
110	10	-3.19583
222	20	-.3221495
336	30	9.017804E-02
451	40	7.951378E-02
566	50	7.312014E-02
797	70	-5.237437E-02
1028	90	-.131265
1257	110	-1.920796E-02
1486	130	5.737073E-02
1716	150	5.306167E-02
1947	170	3.904608E-03
2177	190	8.489426E-03
2408	210	-2.917211E-02
2638	230	-2.250497E-02
2867	250	1.705597E-02

\*\*\*\*\*

SO2 - CHLOROBENZENE

TR = 24.3	TE = -5
VI = 21423.7	V2 = 99.34598
DR = 1.101586	DE = 1.133
PO2 = .234514	PA = 101.6316
FR = 1.464333E-04	R = 0
S = 8.578811E-02	VO = .5754347
L = 602.9215	X = .7365556

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
113	10	-2.624192
227	20	-.2460704
343	30	-2.517637E-03
459	40	.1197053
576	50	2.122948E-02
693	60	-4.430714E-02
810	80	-1.902218E-02
1160	100	-8.955748E-02
1392	120	6.269213E-03
1625	140	1.349495E-02
1858	160	1.891494E-02
2091	180	2.313089E-02
2325	200	-1.639288E-02
2557	220	2.926382E-02
2791	240	-4.189634E-03
2908	250	-1.889901E-02

\*\*\*\*\*

SO2 - DIMETHYLDIBROMIDE

TR = 24.5	TE = 60
VI = 27021.1	V2 = 80.14471
DR = .94474	DE = .9121
PO2 = 3.411986	PA = 100.3917
FR = 2.923334E-04	R = 0
S = 6.330489E-02	VO = .6115665
L = 213.598	X = .3878298

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
152	10	-2.285633
309	20	-.8564868
622	40	3.199647E-02
780	50	2.125238E-02
938	60	1.409093E-02
1255	80	-7.394088E-02
1570	100	-2.365107E-04
1886	120	-3.814552E-03
2201	140	3.885961E-02
2517	160	3.128076E-02
2834	180	-9.790098E-03
2992	190	-1.041505E-02
3150	200	-1.097749E-02

SO2 - DIMETHYLFORMAMIDE

TR = 25.4	TE = 60
V1 = 27021.1	V2 = 80.14471
DR = .943948	DE = .9121
PO2 = 3.451983	PA = 100.7384
FR = 2.923334E-04	R = 0
S = 6.326595E-02	VO = -.1164497
L = 213.6736	X = .3879138

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
158	10	1.21898
317	20	.3066665
475	30	.2175604
633	40	.1730666
792	50	1.965718E-02
950	60	2.300792E-02
1109	70	-6.492756E-02
1267	80	-5.185298E-02
1425	90	-4.167503E-02
1583	100	-3.352757E-02
1740	110	3.065182E-02
2057	130	-1.660505E-02
2372	150	3.310218E-02
2689	170	-3.329076E-03
3005	190	1.212687E-03
3163	200	3.14341E-03

\*\*\*\*\*

SO2 - DIMETHYLFORMAMIDE

TR = 25	TE = 60
V1 = 27021.1	V2 = 80.14471
DR = .9443	DE = .9121
PO2 = 3.411986	PA = 100.3251
FR = 2.923334E-04	R = 0
S = 6.334045E-02	VO = -.3792005
L = 213.8203	X = .3800768

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
165	10	-.7145854
322	20	-8.204399E-02
480	30	-.0806539
638	40	-7.994931E-02
796	50	-7.952655E-02
954	60	-7.925106E-02
1111	70	1.137999E-02
1269	80	2.193456E-04
1426	90	6.195514E-02
1584	100	4.795848E-02
1742	110	3.650957E-02
2058	130	1.890099E-02
2374	150	5.981803E-03
2690	170	-3.877377E-03
3006	190	-1.167563E-02

## S02 - DIMETHYLFORMAMIDE

TR = 24.7	TE = -5
V1 = 21423.7	V2 = 74.45508
DR = .944564	DE = .9818
P02 = 0	PA = 100.2718
FR = 1.464333E-04	R = 0
S = .2481725	VO = -.4751921
L = 1761.592	X = .8595933

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
83	20	-.6118304
123	30	-.1664502
162	40	.6827688
203	50	.192746
243	60	.2829461
325	80	-.2255491
405	100	-3.464071E-02
486	120	-.1137138
567	140	-.1701289
646	160	9.746506E-02
726	180	.1680608
808	200	-2.406495E-02
889	220	-6.818801E-02
969	240	-1.627578E-03
1009	250	.027681

\*\*\*\*\*

## S02 - DIMETHYLFORMAMIDE

TR = 24.3	TE = -5
V1 = 21423.7	V2 = 74.45508
DR = .944916	DE = .9818
P02 = 0	PA = 100.3251
FR = 1.464333E-04	R = 0
S = .2481471	VO = 2.386404
L = 1760.754	X = .8595358

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
76	20	-5.640719
154	40	-1.358928
233	60	-.2571344
272	70	.239974
313	80	-8.057899E-03
394	100	-.106256
474	120	3.489758E-02
555	140	-4.145402E-02
635	160	5.638439E-02
717	180	-.1430666
797	200	-5.480241E-02
877	220	1.753681E-02
957	240	7.789317E-02
998	250	5.108904E-03

\*\*\*\*\*

SO2 - DIMETHYLFORMAMIDE

TR = 24.1	TE = -5
V1 = 21423.7	V2 = 74.45508
DR = .945092	DE = .9818
P02 = 0	PA = 100.2584
FR = 1.464333E-04	R = 0
S = .2478592	VO = 3.337219
L = 1758.384	X = .8593731

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
74	20	-7.743965
112	30	-3.529057
151	40	-1.874092
230	60	-.5714318
312	80	-.8296604
350	90	-9.760328E-02
390	100	-2.296395E-03
471	120	-6.570311E-02
550	140	.243621
629	160	.4768592
709	180	.5196971
789	200	.5539936
877	220	-.3215612
957	240	-.2238528
998	250	-.2794901

\*\*\*\*\*

SO2 - DIMETHYLFORMAMIDE

TR = 24	TE = 25
V1 = 24047.26	V2 = 77.41184
DR = .94518	DE = .9443
P02 = 0	PA = 101.4317
FR = 1.464333E-04	R = 0
S = .0648446	VO = .41476
L = 442.4143	X = .5874928

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
152	10	-2.639807
304	20	-.6335491
456	30	5.371434E-02
608	40	.4009139
761	50	.4792913
914	60	.5316109
1220	80	.5970919
1555	100	-1.232715
1850	120	-.3133903
2156	140	-.1566908
2462	160	-3.882807E-02
2767	180	8.910588E-02
2920	190	.1259632
3073	200	.1591503

## SD2-DIMETHYLFORMAMIDE

TR = 23.9	TE = 25
V1 = 24047.26	V2 = 77.41184
DR = .945268	DE = .9443
PO2 = 0	PA = 99.61847
FR = 1.464333E-04	R = 0
S = 6.596686E-02	VO = 2.077667
L = 450.0294	X = .5916225

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
145	10	-14.11047
293	20	-6.568061
443	30	-4.156372
748	50	-2.763234
1053	70	-2.153697
1358	90	-1.811748
1664	110	-1.65094
1970	130	-1.539304
2275	170	11.73018
2578	170	-1.243305
2883	190	-1.175552
3188	210	-1.120642
3494	230	-1.103288

\*\*\*\*\*

## SD2 - DIMETHYLFORMAMIDE

TR = 22.9	TE = 25
V1 = 24047.26	V2 = 77.41184
DR = .946148	DE = .9443
PO2 = 0	PA = 100.0584
FR = 1.464333E-04	R = 0
S = 6.633925E-02	VO = -1.56019
L = 452.149	X = .5927572

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
171	10	2.209555
322	20	1.004763
472	30	.8337809
624	40	.4129416
776	50	.1621277
928	60	-4.38671E-03
1080	70	-.122987
1231	80	-.1291132
1383	90	-.2073346
1536	100	-.33576
1678	110	.221339
1830	120	.1329853
1982	130	5.834608E-02
2134	140	-5.536454E-03

\*\*\*\*\*

## SH2 - CHLOROBENZENE

TR = 23.5	TE = 25
V1 = 24264.5	V2 = 102.2427
DR = 1.10237	DE = 1.1009
PO2 = 1.595602	PA = 101.365
FR = .01754	R = 0
S = .2188589	VO = -5.916585E-02
L = 12.65018	X = 5.060625E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.05	1.436656
10	2.1	-1.381747
15	3.19	-1.045917
20	4.27	-1.111906
25	5.39	-.412143
30	6.49	-.2551422
35	7.59	-.1433417
0	0	-100-
45	9.79	5.270327E-03
50	10.89	5.716935E-02
55	11.98	1.603291E-02
60	13.08	5.838465E-02
65	14.18	9.415124E-02
70	15.27	5.926033E-02

\*\*\*\*\*

## SH2 - CHLOROBENZENE

TR = 22.5	TE = 25
V1 = 24264.5	V2 = 102.2427
DR = 1.10335	DE = 1.1009
PO2 = 1.595602	PA = 100.1384
FR = .01754	R = 0
S = .2188365	VO = -.3621818
L = 12.64005	X = 5.056776E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1	36.61187
10	1.95	6.780075
15	2.95	1.014744
20	4	-.3623895
25	5.1	-.1707106
30	6.2	-4.696746E-02
35	7.3	3.979581E-02
40	8.399999	.1039222
45	9.5	.1332742
50	10.6	.1924089
55	11.7	.2241993
60	12.75	-.1410493
65	13.84	-.1600901
70	14.95	-4.261969E-02

\*\*\*\*\*

SH2 -HEXANE

TR = 23	TE = 25
V1 = 24264.5	V2 = 131.6901
DR = .6564	DE = .6544
PO2 = 20.16901	PA = 100.805
FR = .01754	R = 0
S = .1004285	VO = 2.642857E-02
L = 7.08348	X = 3.702076E-02

TIME IN MIN.	VOL. IN-CC	DEV. IN (%)
0	0	0
5	.52	-1.621547
10	1.03	-6.922075E-02
15	1.53	-.1863199
20	2.02	-.7370128
25	2.54	.1126905
30	3.05	.3526054
35	3.55	.2421129
40	4.05	.1590688
45	4.54	-.1256366
50	5.04	-.1555811

\*\*\*\*\*

SH2 -HEXANE

TR = 22.4	TE = 25
V1 = 24264.5	V2 = 131.6901
DR = .657	DE = .6544
PO2 = 20.16901	PA = 101.085
FR = .01754	R = 0
S = 9.930211E-02	VO = 3.641505E-02
L = 6.993313	X = 3.656675E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
2	.24	2.119285
5	.54	1.327468
8.5	.87	-1.190586
10	1.03	5.477367E-02
15	1.53	.2656289
20	2.02	-.1214929
24	2.42	1.382428E-02
30	3.01	-.1816675
35	3.51	-5.663138E-02
40	4.01	3.744752E-02
45	4.51	.1107784
50	5	-.0303939

\*\*\*\*\*

SH2 -HEXANE

TR =	24.8	TE =	-5
V1 =	21723.19	V2 =	126.2312
DR =	.6546	DE =	.6827
PO2 =	4.594289	PA =	100.0851
FR =	.000734	R =	0
S =	1.219597E-02	VO =	-8.355372E-03
L =	18.0845	X =	9.509396E-02

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
410	5	.1604861
822	10	-.1670142
1232	15	-.1137075
1641	20	-2.611433E-02
2050	25	2.651154E-02
2460	30	2.093435E-02
2871	35	-1.789318E-02
3281	40	-1.651491E-02
3690	45	.0116489

\*\*\*\*\*

SH2 -HEXANE

TR =	23.1	TE =	-5
V1 =	21723.19	V2 =	126.2312
DR =	.6563	DE =	.6827
PO2 =	4.594289	PA =	100.5384
FR =	.000734	R =	0
S =	1.211743E-02	VO =	1.088292E-03
L =	17.9182	X =	9.430198E-02

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
414	5	-.3528404
826	10	-.100759
1238	15	-.016445
1650	20	2.575584E-02
2063	25	2.609371E-03
2476	30	-.0128221
2889	35	-2.383073E-02
3301	40	-1.811949E-03
3713	45	1.532048E-02

\*\*\*\*\*

## SH2 - HEXANE

TR = 25.2	TE = -5
V1 = 21723.19	V2 = 126.2312
DR = .6542	DE = .6827
P02 = 4.594289	FA = 100.1651
FR = .000734	R = 0
S = .0120973	VO = 4.698386E-02
L = 17.94898	X = 9.444863E-02

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
413	5	-.9360088
827	10	-.5118926
1239	15	-.2363835
1651	20	-9.805591E-02
2061	25	8.196869E-02
2475	30	4.065576E-02
2890	35	-2.339497E-02
3303	40	-1.093745E-02
3716	45	-1.246119E-03

\*\*\*\*\*

## SH2 - ACETONITRILE

TR = 24.1	TE = -5
V1 = 21723.19	V2 = 50.67021
DR = .77893	DE = .8102
P02 = 2.277146	FA = 100.2851
FR = .000734	R = 0
S = 3.375715E-02	VO = 5.542417
L = 48.84208	X = .1022744

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
141	5	-51.46656
286	15	-1.296058
432	20	-.6236139
576	25	5.389256E-02
725	30	-5.446963E-02
1020	40	6.326861E-02
1317	50	-1.167284E-03
1614	60	-4.407211E-02
1910	70	-2.653235E-02
2205	80	2.883789E-02
2502	90	-3.119477E-03
2798	100	5.073805E-03
3095	110	-1.890345E-02
3390	120	1.737261E-02
3687	130	-3.873235E-03
3983	140	2.03818E-03

\*\*\*\*\*

SH2 - ACETONITRILE

TR = 24.5  
 V1 = 21723.19  
 DR = .77815  
 P02= 2.277146  
 FR = .000734  
 S = .034002  
 L = 49.24519

TE = -5  
 V2 = 50.67021  
 DE = .8102  
 PA = 99.72512  
 R = 0  
 VD = .4349062  
 X = .1030315

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
141	5	-4.382869
284	10	-.9064536
429	15	-.1448881
576	20	-.1001974
724	25	-.2089816
870	30	-5.546035E-02
1164	40	-8.308155E-02
1457	50	.0483785
1751	60	4.600097E-02
2044	70	9.294694E-02
2339	80	4.302922E-02
2633	90	4.204716E-02
2928	100	7.233109E-03
3223	110	-2.123295E-02
3517	120	-1.662296E-02
3812	130	-3.085959E-02

\*\*\*\*\*

SH2 - ACETONITRILE

TR = 24.4  
 V1 = 21723.19  
 DR = .77832  
 P02= 2.277146  
 FR = .000734  
 S = 3.387801E-02  
 L = 49.05489

TE = -5  
 V2 = 50.67021  
 DE = .8102  
 PA = 99.73846  
 R = 0  
 VD = .5182057  
 X = .1026743

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
139	5	-4.347395
279	10	.2991923
427	15	.1060083
575	20	9.699808E-03
723	25	-4.800399E-02
871	30	-8.642986E-02
1166	40	-4.989041E-02
1460	50	3.981076E-02
1755	60	4.316285E-02
2052	70	-5.123237E-02
2346	80	4.978428E-03
2639	90	.0864394
2935	100	4.986006E-02
3232	110	-1.084644E-02
3528	120	-3.317686E-02
3823	130	-2.601534E-02

## SH2 - ACETONITRILE

TR = 25	TE = 25
V1 = 24264.5	V2 = 52.81488
DR = .7773	DE = .7773
P02= 11.50972	PA = 100.0051
FR = .01754	R = 0
S = .3481818	VO = .2863636
L = 22.31462	X = 4.632087E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.8	-11.21076
15	3.6	-34.65347
15	5.3	-3.795376
20	7.1	-2.068967
25	8.899999	-1.011122
30	10.7	-.2964865
35	<del>12.5</del>	.2186625
40	14.3	.6076122
45	16	.284903
50	17.8	.5908048
55	19.6	.8419124
60	21.4	1.051726
65	23.2	1.22968
70	24	-2.672808

\*\*\*\*\*

## SH2 - ACETONITRILE

TR = 21.8	TE = 25
V1 = 24264.5	V2 = 52.81488
DR = .78274	DE = .7773
P02= 11.50972	PA = 101.3517
FR = .01754	R = 0
S = .3578182	VO = -.2509091
L = 22.73128	X = 4.714499E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.75	13.77069
10	3.35	.6830595
15	5.15	.6574312
20	6.9	-7.898886E-02
25	8.7	6.272971E-02
30	10.48	-3.469515E-02
35	12.26	-.1037075
40	14.06	-1.292652E-02
45	15.87	.1204389
50	17.64	0
55	19.44	5.615309E-02
60	21.22	8.566725E-03
65	23.01	1.185498E-02
70	24.78	-.0659901

\*\*\*\*\*

## SH2 - ACETONITRILE

TR = 23.4  
 V1 = 27180.7  
 DR = .78002  
 PQ2 = 49.45326  
 FR = .01754  
 S = .1000334  
 L = 10.38137

TE = 60  
 V2 = 55.54459  
 DE = .7391  
 PA = 102.0183  
 R = 0  
 VO = 5.310648E-02  
 X = 2.077394E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	.55	-.5916488
10	1.05	1.3266078
15	1.55	-.2322176
20	2.05	-.1838021
25	2.55	-.1543501
30	3.06	.1928826
35	3.56	.1610374
40	4.06	.137061
45	4.55	-.1012175
50	5.05	-9.450376E-02
55	5.55	-8.899906E-02
60	6.06	8.073401E-02

\*\*\*\*\*

## SH2 - ACETONITRILE

TR = 25  
 V1 = 27180.7  
 DR = .7773  
 PQ2 = 49.45326  
 FR = .01754  
 S = .1000333  
 L = 10.48828

TE = 60  
 V2 = 55.54459  
 DE = .7391  
 PA = 101.285  
 R = 0  
 VO = -2.222222E-04  
 X = 2.098338E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	.45	-9.989942
10	1	-1.104948E-02
15	1.5	-1.845814E-02
20	2	-.0221561
25	2.5	-2.437951E-02
30	3.01	.3073851
35	3.5	-2.692043E-02
40	4	-2.770848E-02
45	4.49	-.2504911
50	5	-2.882126E-02
55	5.5	-2.922591E-02
60	6.01	.1370581

\*\*\*\*\*

SH2 - ETHYLACETATE

TR = 24.1	TE = 60
V1 = 27180.7	V2 = 103.5822
DR = .89528	DE = .8506
P02 = 56.04206	PA = 100.765
FR = .01754	R = 0
S = 9.247617E-02	VO = 6.714286E-02
L = 11.03421	X = 4.035311E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	.5	-5.575515
10	.98	-1.200164
15	1.47	1.080575
20	1.94	1.217422
25	2.4	.8807282
30	2.81	-1.106067
35	3.28	-.7206446
40	3.75	-.4298662
45	4.23	3.380712E-02
50	4.7	.1928918
55	5.17	.3234382

\*\*\*\*\*

SH2 - ETHYLACETATE

TR = 23	TE = 60
V1 = 27180.7	V2 = 103.5822
DR = .8966	DE = .8506
P02 = 56.04206	PA = 100.8184
FR = .01754	R = 0
S = 9.283331E-02	VO = .0475
L = 11.05287	X = 4.041861E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	.5	-2.280105
10	.97	-.5977548
15	1.45	.6944688
20	1.91	.3063674
25	2.37	.0703981
30	2.83	-8.823804E-02
35	3.29	-.2021954
40	3.76	-2.213125E-02
45	4.22	-.1183234
50	4.69	.0177956
55	5.16	.1293845

\*\*\*\*\*

## SH2 --ETHYLACETATE

TR = 22.8	TE = 25
V1 = 24264.4	V2 = 98.53164
DR = .8968401	DE = .8942
P02= 12.90694	PA = 101.765
FR = .01754	R = 0
S = .36	VO = 0
L = 23.17982	X = 8.602968E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5.5	1.9	-4.040406
10	3.6	-6.622738E-06
15	5.4	0
20	7.2	-6.622738E-06
25	9	0
30	10.8	0
35	12.6	0
40	14.4	-6.622738E-06
45	16.2	0
50	18	0
55	19.8	-9.633073E-06
60	21.6	0
65	23.4	-8.151063E-06
70	25.2	0

\*\*\*\*\*

## SH2 -- ETHYLACETATE

TR = 22.6	TE = 25
V1 = 24264.5	V2 = 98.53164
DR = .89708	DE = .8942
P02= 12.90694	PA = 100.805
FR = .01754	R = 0
S = .357	VO = -.2331736
L = 23.01401	X = 8.546657E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.95	25.65838
10	3.6	7.886943
15	5.2	1.526275
20	6.9	-9.883556E-02
25	8.7	9.403074E-02
30	10.5	.2211866
35	12.3	.3113213
40	14	-.333359
45	15.8	-.2010196
50	17.6	-9.551453E-02
55	19.4	-9.417876E-03
60	21.2	6.218042E-02
65	23	.1226435
70	24.75	-2.757381E-02

SH2 - ETHYLACETATE

TR = 23.5	TE = -5
V1 = 21723.19	V2 = 94.43408
DR = .896	DE = .933
PO2 = 2.487795	FA = 101.725
FR = .000734	R = 0
S = 3.615402E-02	VO = -4.371133E-02
L = 52.33676	X = .1853467

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
139	5	.3674033
278	10	-7.100782E-02
417	15	-.216295
556	20	-.108734
831	30	-9.282344E-04
1108	40	-3.734148E-02
1384	50	1.310139E-02
1660	60	4.675189E-02
1937	70	1.912074E-02
2214	80	-1.602147E-03
2490	90	2.245242E-02
2767	100	5.539248E-03
3044	110	-8.294545E-03
3321	120	-1.981978E-02
3597	130	-1.760599E-03

\*\*\*\*\*

SH2 - ETHYLACETATE

TR = 24.8	TE = -5
V1 = 21723.19	V2 = 94.43408
DR = .8944401	DE = .933
PO2 = 2.487795	FA = 101.5783
FR = .000734	R = 0
S = 3.658472E-02	VO = .4225068
L = 53.05172	X = .1874042

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
138	5	-8.612346
287	10	-8.444379
423	15	-5.647585
562	20	-4.685299
700	25	-3.963661
978	35	-3.321246
1255	45	-2.883988
1532	55	-2.603681
1808	65	-2.355029
2085	75	-2.218535
2361	85	-2.072652
2638	95	-1.994167
2914	105	23.4139
3191	115	-1.847285
3468	125	-1.805469
3745	135	-1.769815

## SH2 - ETHYLACETATE

TR = 24	TE = -5
V1 = 21723.19	V2 = 94.43408
DR = .8954001	DE = .933
P02 = 2.487795	PA = 101.3917
FR = .000734	R = 0
S = 3.581485E-02	VO = .4498368
L = 51.88652	X = .1840456

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
135	5	-5.389784
272	10	-1.878794
410	15	-.8849352
547	20	-.2023884
686	25	-7.524546E-02
826	30	-.1095587
1103	40	.1160919
1384	50	-3.516677E-02
1664	60	-7.619135E-02
1942	70	-3.258735E-03
2219	80	9.635864E-02
2498	90	9.415901E-02
2777	100	9.240862E-02
3056	110	9.096958E-02
3343	120	-.1488501
3619	130	-4.903882E-02

\*\*\*\*\*

## SH2 - DIMETHYLFORMANIDE

TR = 25	TE = 60
V1 = 27180.7	V2 = 80.14471
DR = .9443	DE = .9121
P02 = 3.451983	PA = 100.6584
FR = .01754	R = 0
S = .3381818	VO = -.1390909
L = 19.24922	X = 5.370958E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.63	5.038073
10	3.25	.2242779
15	4.9	-.6817708
20	6.6	-.3705261
25	8.3	-.1858502
30	10	-.0635887
35	11.7	2.330932E-02
40	13.4	8.827138E-02
45	15.1	.1386706
50	16.79	.1192633
55	18.5	.2117508
60	20.2	.2391022
65	21.79	-.2413927
70	23.48	-.1854211

SH2 - DIMETHYLFORMAMIDE

TR = 24.7	TE = 60
V1 = 27180.7	V2 = 80.14471
DR = .944564	DE = .9121
P02 = 3.451983	PA = 101.0317
FR = .01754	R = 0
S = .3427276	VO = -.3727355
L = 19.49941	X = 5.436971E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.65	23.05148
10	3.2	4.762096
15	4.8	.6673971
20	6.5	.2805489
25	8.2	5.548344E-02
30	9.899999	-9.173823E-02
35	11.6	-.1955476
40	13.3	-.2726793
45	15.05	-1.901012E-05
50	16.8	.2168853
55	18.5	.122974
60	20.2	4.499401E-02
65	21.9	-2.079362E-02
70	23.6	-.0770266

\*\*\*\*\*  
 SH2 - DIMETHYLFORMAMIDE

TR = 24.8	TE = 60
V1 = 27180.7	V2 = 80.14471
DR = .944476	DE = .9121
P02 = 3.451983	PA = 100.8717
FR = .01754	R = 0
S = .3410364	VO = -.6393636
L = 19.40609	X = .0541236

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.55	45.42815
10	3.05	10.06855
15	4.55	1.649123
20	6.2	.3014751
25	7.9	.1705878
30	9.600001	8.623281E-02
35	11.3	2.734333E-02
40	13	-1.609983E-02
45	14.68	-.185453
50	16.39	-.1368298
55	18.09	-.1525446
60	19.8	-.115127
65	21.5	-.1300715
70	23.32	.3736676

\*\*\*\*\*  
 SH2 - DIMETHYLFORMAMIDE

## SH2 - DIMETHYLFORMAMIDE

IR = 23.6	TE = 25
V1 = 24264.5	V2 = 77.41184
DR = .945532	DE = .9443
PO2 = 0	FA = 102.0583
FR = .01754	R = 0
S = .7156	VO = .06
L = 40.74507	X = .1150367

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	3.7	1.704234
10	7.25	.4711741
15	10.8	5.558241E-02
20	14.4	.1948158
25	18	.2785579
30	21.6	.3344505
35	25.2	.374412
40	28.25	-1.513039
45	32.32	.1797739
50	35.89	.1395068
55	39.43	.0304359
60	43.04	.1028321
65	46.6	5.581904E-02

\*\*\*\*\*

## SH2 - DIMETHYLFORMAMIDE

TR = 22.9	TE = 25
V1 = 24264.5	V2 = 77.41184
DR = .946148	DE = .9443
PO2 = 0	FA = 101.9116
FR = .01754	R = 0
S = .7247273	VO = -.2109091
L = 41.2379	X = .1162663

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	3.5	2.557277
10	7.1	.9043919
15	10.7	.3752342
20	14.3	.1145653
25	18	.5178213
30	21.6	.3208959
35	25.2	.1807066
40	28.3	-1.661613
45	32.5	.3030165
50	36.1	.2069282
55	39.7	.1284038
60	43.3	6.303067E-02
65	46.9	7.760136E-03

\*\*\*\*\*

## SH2 - DIMETHYLFORMAMIDE

TR = 26.4	TE = -5
V1 = 21723.19	V2 = 74.45508
DR = .9430679	DE = .9818
P02 = 0	PA = 100.7117
FR = .000734	R = 0
S = 6.890005E-02	VD = 1.010825
L = 97.72456	X = .2509059

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
137	10	-4.307424
280	20	-1.491601
422	30	-.2879917
567	40	-.1925093
714	50	-.4092376
859	60	-.3255511
1145	80	.1234241
1435	100	.1177456
1725	120	.1139539
2014	140	.1606056
2306	160	.0660856
2596	180	6.929746E-02
2885	200	.1063829
3184	220	-.1763091
3616	250	-6.132102E-02

\*\*\*\*\*

## SH2 - DIMETHYLFORMAMIDE

TR = 24.5	TE = -5
V1 = 21723.19	V2 = 74.45508
DR = .94474	DE = .9818
P02 = 0	PA = 100.765
FR = .000734	R = 0
S = 6.869208E-02	VD = 1.273845
L = 97.25685	X = .2500053

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
139	10	-7.596018
277	20	-1.485363
417	30	.2725955
560	40	.6506833
705	50	.6000528
1004	70	-.3426677
1293	90	-.1028993
1587	110	-.2612881
1876	130	-.1077283
2165	150	5.198431E-03
2455	170	5.125989E-02
2749	190	-5.701126E-02
3039	210	-1.384725E-02
3329	230	2.183811E-02
3618	250	7.934146E-02

## SH2 - DIMETHYLFORMAMIDE

TR = 26.8	TE = -5
V1 = 21723.19	V2 = 74.45508
DR = .942716	DE = .9818
PO2 = 0	PA = 100.3917
FR = .000734	R = 0
S = 6.902542E-02	VO = .4011233
L = 97.93861	X = .2513173

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
140	10	-.6426632
282	20	.6730444
424	30	1.119391
570	40	.6400376
716	50	.3546084
1024	70	-1.523783
1295	90	.2349512
1585	110	.1762996
1874	130	.1890021
2165	150	.1060083
2457	170	2.01061E-03
2746	190	2.899207E-02
3038	210	4.775099E-02
3327	230	-2.116541E-02
3617	250	-2.642124E-02

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## SH2 - DIMETHYLACETAMIDE

TR = 25	TE = -5
V1 = 21723.19	V2 = 89.66859
DR = .9368	DE = .9716
PO2 = 0	PA = 100.0451
FR = .000734	R = 0
S = .0734484	VO = 1.005927
L = 103.7831	X = .2999131

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
130	10	-5.251155
264	20	-1.943015
398	30	-.7883623
532	40	-.2007832
667	50	7.980984E-03
940	70	-6.769220E-02
1211	90	.0534343
1483	110	6.376689E-02
1755	130	7.093323E-02
2028	150	2.715785E-02
2300	170	3.693991E-02
2575	190	-.071288
2847	210	-5.402324E-02
3118	230	-7.834443E-03
3389	250	3.099715E-02

\*\*\*\*\*

## SH2 - DIMETHYLACETAMIDE

TR = 24.5	TE = -5
V1 = 21723.19	V2 = 89.66859
DR = .93733	DE = .9716
PO2 = 0	PA = 101.3517
FR = .000734	R = 0
S = 7.415845E-02	VO = -.8773879
L = 104.7272	X = .3018178

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
139	10	6.037393
273	20	3.263826
416	30	9.166847E-02
551	40	4.023263E-02
686	50	9.392667E-03
857	70	-.131608
1225	90	3.700778E-02
1495	110	9.558464E-03
1764	130	.0476301
2034	150	2.608916E-02
2305	170	-3.400656E-02
2574	190	-3.388943E-03
2844	210	-1.391989E-02
3113	230	9.633868E-03
3383	250	-2.56347E-04

\*\*\*\*\*

## SH2 - DIMETHYLACETAMIDE

TR = 25.5	TE = -5
V1 = 21723.19	V2 = 89.66859
DR = .93627	DE = .9716
PO2 = 0	PA = 101.4717
FR = .000734	R = 0
S = 7.376336E-02	VO = .3878813
L = 104.2872	X = .3009313

TIME IN SEC.	VOL. IN CC	DEV. IN (%)
0	0	0
132	10	-1.231107
266	20	-4.465016E-02
402	30	-.1356571
537	40	2.98509E-03
673	50	-6.121129E-02
844	70	-2.926651E-02
1216	90	-9.338974E-02
1485	110	6.689208E-02
1757	130	7.606498E-03
2027	150	6.255946E-02
2299	170	1.773924E-02
2570	190	2.120618E-02
2842	210	-1.112315E-02
3114	230	-3.780759E-02
3384	250	-1.238998E-03

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## SH2 - DIMETHYLACETAMIDE

TR = 23	TE = 25
V1 = 24264.5	V2 = 92.99958
DR = .93892	DE = .9368
P02= 0	PA = 101.4317
FR = .01754	R = 0
S = .7821673	VO = -.1089259
L = 44.49258	X = .1456851

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	3.9	2.580011
10	7.8	1.13129
15	11.6	-.2028846
20	15.5	-.2215665
25	19.45	2.440433E-02
30	23.35	-2.607527E-02
35	27.3	.121288
40	31.2	7.132576E-02
45	35.1	3.248438E-02
50	39	1.44765E-03
55	42.9	-2.393172E-02
60	46.8	-.0450795

\*\*\*\*\*

## SH2 - DIMETHYLACETAMIDE

TR = 22.8	TE = 25
V1 = 24264.5	V2 = 92.99958
DR = .939132	DE = .9368
P02= 0	PA = 100.3651
FR = .01754	R = 0
S = .783	VO = -.07
L = 44.52996	X = .1457897

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	3.9	1.430437
10	7.8	.5154696
15	11.7	.2141377
20	15.6	6.414515E-02
25	19.5	-2.563993E-02
30	23.4	-8.539904E-02
35	27.35	5.487946E-02
40	31.25	0
45	35.15	-4.265432E-02
50	39.1	5.116848E-02
55	43	1.163174E-02
60	46.9	-2.131384E-02

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SH2 - DIMETHYLACETAMIDE

TR = 24	TE = 60
V1 = 27180.7	V2 = 96.17177
DR = .93786	DE = .9059
P02 = 2.125692	PA = 100.9517
FR = .01754	R = 0
S = .3565455	VO = -.2081818
L = 20.02931	X = 6.617849E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.7	7.967675
10	3.35	-.2166259
15	5.1	-.7782095
20	6.9	-.3282957
25	8.7	-6.265111E-02
30	10.5	.1126786
35	12.3	.2370797
40	14.05	-2.587487E-02
45	15.85	8.611537E-02
50	17.6	-1.1083521
55	19.4	-9.378556E-03
60	21.2	7.295529E-02
65	22.95	-7.519849E-02
70	24.75	0

\*\*\*\*\*

SH2 - DIMETHYLACETAMIDE

TR = 23.8	TE = 60
V1 = 27180.7	V2 = 96.17177
DR = .938072	DE = .9059
P02 = 2.125692	PA = 100.6451
FR = .01754	R = 0
S = .3591818	VO = -.2540827
L = 20.17398	X = 6.662465E-02

TIME IN MIN.	VOL. IN CC	DEV. IN (%)
0	0	0
5	1.7	10.25884
10	3.4	1.86547
15	5.15	.3185946
20	6.95	.2950525
25	8.75	.2812124
30	10.5	-.2031279
35	12.3	-.1402954
40	14.1	-9.346058E-02
45	15.9	-5.719373E-02
50	17.7	-2.828972E-02
55	19.5	+4.714352E-03
60	21.3	1.490282E-02
65	23.1	3.146051E-02
70	24.9	4.562099E-02

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